

**3 • Molecules & Compounds****Mole Calculations - Difficulty Level 1**

$$1 \text{ mole} = 6.02 \times 10^{23} \text{ molecules} = 22.4 \text{ L (@ STP)}$$

1. Calculate the mass of 1.58 moles
- $\text{CH}_4$
- . [molar mass
- $\text{CH}_4 = 16.0 \text{ g/mol}$
- ]

G: 1.58 moles  $\text{CH}_4$ D: ? g  $\text{CH}_4$ 

$$\frac{1.58 \text{ moles } \text{CH}_4}{1} \times \frac{16.0 \text{ g } \text{CH}_4}{1 \text{ mole } \text{CH}_4} = 25.28 \approx \boxed{25.3 \text{ g } \text{CH}_4}$$

2. What volume will 7.29 moles of
- $\text{CO}_2$
- gas occupy at STP?

G: 7.29 moles  $\text{CO}_2$ D: ? L  $\text{CO}_2$ 

$$\frac{7.29 \text{ moles } \text{CO}_2}{1} \times \frac{22.4 \text{ L } \text{CO}_2}{1 \text{ mole } \text{CO}_2} = 163.376 \approx \boxed{163 \text{ L } \text{CO}_2}$$

3. How many molecules are there in a 0.00583 mole sample of
- $\text{H}_2\text{O}$
- ?

G: 0.00583 moles  $\text{H}_2\text{O}$ D: ? molecules  $\text{H}_2\text{O}$ 

$$\frac{0.00583 \text{ moles } \text{H}_2\text{O}}{1} \times \frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mole } \text{H}_2\text{O}} = \boxed{3.51 \times 10^{21} \text{ molecules}}$$

4. What mass of
- $\text{CO}_2$
- gas occupies a volume of 100. Liters at STP? [molar mass
- $\text{CO}_2 = 44.0 \text{ g/mol}$
- ]

G: 100. Liters  $\text{CO}_2$ D: ? g  $\text{CO}_2$ 

$$\frac{100. \text{ Liters } \text{CO}_2}{22.4 \text{ L } \text{CO}_2} \times \frac{44.0 \text{ g } \text{CO}_2}{1 \text{ mole } \text{CO}_2} = \boxed{196 \text{ g } \text{CO}_2}$$

5. How many molecules are in a 35.0 gram sample of
- $\text{H}_2\text{O}$
- ? [molar mass
- $\text{H}_2\text{O} = 18.0 \text{ g/mol}$
- ]

G: 35.0 g  $\text{H}_2\text{O}$ D: ? molecules  $\text{H}_2\text{O}$ 

$$\frac{35.0 \text{ g } \text{H}_2\text{O}}{18.0 \text{ g } \text{H}_2\text{O}} \times \frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mole } \text{H}_2\text{O}} = \boxed{1.17 \times 10^{24} \text{ molecules}}$$

6. What volume will
- $5.25 \times 10^{22}$
- molecules of
- $\text{CH}_4$
- occupy at STP?

G:  $5.25 \times 10^{22}$  molecules  $\text{CH}_4$ 

D: ? L

$$\frac{5.25 \times 10^{22} \text{ molecules } \text{CH}_4}{6.02 \times 10^{23} \text{ molecules}} \times \frac{22.4 \text{ L } \text{CH}_4}{1 \text{ mole } \text{CH}_4} = \boxed{1.95 \text{ L } \text{CH}_4}$$

# Moles And Particles Answers

**Camilla Rothe**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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