

Module 4 Practice Problems

1. Nitric acid is produced via the steps for ammonia synthesis. How many moles of H_2 are used to produce 100.00 moles of ammonia (nitrogen)?



$$\frac{1 \text{ mol } \text{H}_2}{1.5 \text{ mol } \text{CO}_2} \times 100 \text{ mol } \text{CO}_2 \times \left(\frac{2 \text{ mol } \text{H}_2}{1 \text{ mol } \text{CO}_2} \right) = \boxed{133.33 \text{ mol } \text{CO}_2}$$

2. How many moles of nitrogen gas are produced from 100.00 moles of ammonia (nitrogen) and 100.00 moles of oxygen gas?



3. From a very small amount of ammonia, a product is produced by the following reaction:



How many moles of ammonia are produced from 100.00 moles of nitrogen gas and 100.00 moles of hydrogen gas?

$$\frac{2 \text{ mol } \text{NH}_3}{1 \text{ mol } \text{N}_2} \times 100 \text{ mol } \text{N}_2 \times \left(\frac{2 \text{ mol } \text{H}_2}{1 \text{ mol } \text{N}_2} \right) = \boxed{100 \text{ mol } \text{NH}_3}$$

4. The decomposition of ammonia is represented by the following reaction:



$$100 \text{ mol } \text{NH}_3 \times \left(\frac{1 \text{ mol } \text{N}_2}{2 \text{ mol } \text{NH}_3} \right) = \boxed{50 \text{ mol } \text{N}_2}$$

5. The reaction of phosphorus trichloride is represented by the following reaction:



How many moles of phosphorus trichloride are produced from 100.00 moles of hydrogen gas and 100.00 moles of water?

$$100 \text{ mol } \text{H}_2 \times \left(\frac{1 \text{ mol } \text{PCl}_3}{3 \text{ mol } \text{H}_2} \right) = 33.33 \text{ mol } \text{PCl}_3$$

$$100 \text{ mol } \text{H}_2\text{O} \times \left(\frac{1 \text{ mol } \text{PCl}_3}{3 \text{ mol } \text{H}_2\text{O}} \right) = 33.33 \text{ mol } \text{PCl}_3$$

$$\text{Limiting reagent is } \text{H}_2 \text{ so } \boxed{33.33 \text{ mol } \text{PCl}_3}$$

6. From decomposition of calcium carbonate, how many moles of carbon dioxide are produced from 100.00 moles of calcium carbonate?



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The Practice of Chemistry Donald J. Wink, Sharon Fetzner-Gislason, Sheila McNicholas, 2003-03 Students can't do chemistry if they can't do the math. The Practice of Chemistry First Edition is the only preparatory chemistry text to offer students targeted consistent mathematical support to make sure they understand how to use math especially algebra in chemical problem solving. The book's unique focus on actual chemical practice, extensive study tools and integrated media makes The Practice of Chemistry the most effective way to prepare students for the standard general chemistry course and bright futures as science majors. This special PowerPoint tour of the text was created by Don Wink http://www.bfwpub.com/pdfs/wink/POC/PowerPoint_Final.ppt 832KB

The Software Finder, 1983

Ebook: Chemistry Julia Burdge, 2014-10-16

Chemistry Third Edition by Julia Burdge offers a clear writing style written with the students in mind. Julia uses her background of teaching hundreds of general chemistry students per year and creates content to offer more detailed explanation on areas where she knows they have problems. With outstanding art, a consistent problem solving approach, interesting applications woven throughout the chapters and a wide range of end of chapter problems, this is a great third edition text.

Software for Schools, 1987

Holt Chemistry Ralph Thomas Myers, 2004

[The Software Encyclopedia](#), 1986

Jawahar Navodaya Vidyalaya Entrance Class XI Book (English Edition) - 15 Practise Tests (1500 Solved Questions) EduGorilla Prep Experts, Best Selling Book in English Edition for JNV Jawahar Navodaya Vidyalaya Entrance Exam For Class XI with objective type questions as per the latest syllabus. JNV Jawahar Navodaya Vidyalaya Entrance Exam For Class XI Preparation Kit comes with 15 Practise Tests with the best quality content. Increase your chances of selection by 16X. JNV Jawahar Navodaya Vidyalaya Entrance Exam For Class XI Prep Kit comes with well structured and 100% detailed solutions for all the questions. Clear exam with good grades using thoroughly researched content by experts.

Chemistry in the Community American Chemical Society, 2006-01-31 This laboratory based text centres itself around decision making activities where students apply their chemistry knowledge to realistic situations. This fifth edition includes more photographs, new drawings and new design.

Chemical Principles Peter Atkins, Loretta Jones, 2007-08 Written for calculus inclusive general chemistry courses, Chemical Principles helps students develop chemical insight by showing the connections between fundamental chemical ideas and their applications. Unlike other texts, it begins with a detailed picture of the atom, then builds toward chemistry's frontier, continually demonstrating how to solve problems, think about nature and matter, and visualize chemical concepts as working chemists do. Flexibility in level is crucial and is largely established through clearly labeling, separating in boxes the calculus coverage in the text. Instructors have the option of whether to incorporate calculus in the coverage of topics. The multimedia integration of Chemical Principles is more deeply established than any other text for this course. Through the unique eBook, the comprehensive Chemistry Portal, Living Graph icons that connect the text to the Web, and a complete set of animations, students can take full advantage of the wealth of resources available to them to help them

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