

MOLE CONCEPT

CLASS Xth

1. Calculate the relative molecular masses (or molecular weights) of the following compounds :

- i) Copper sulphate crystals, $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$.
- ii) Ammonium sulphate, $(\text{NH}_4)_2\text{SO}_4$.
- iii) Cane sugar, $\text{C}_{12}\text{H}_{22}\text{O}_{11}$.

(Given atomic masses Cu = 63.5, S = 32, O = 16, N = 14, C = 12)

2. An organic compound with vapour density = 94 contains. C = 12.67%, H = 2.13%, and Br = 85.11%. Find the molecular formula. [Atomic mass : C = 12, H = 1, Br = 80]

3. (i) The compound A has the following percentage composition by mass. Carbon = 26.7%, oxygen = 71.1%, hydrogen = 2.2%. Determine the Empirical formula of compound A [work to one decimal place] [H = 1, C = 12, O = 16].

ii) If the relative molecular mass of A is 90, what is the molecular formula of A?

(iii) The compound A is weak acid. What is meant by this statement?

4. Determine the empirical formula of the compound whose composition by mass is: 42% nitrogen, 48% oxygen and 9% hydrogen. [H = 1; N = 14; O = 16]

5. Determine the empirical formula of a compound containing 47.9% potassium, 5.5% beryllium and 46.6% fluorine by mass. (Atomic weight of Be = 9; F = 19; K = 39) Work to one decimal place.

6. The Empirical formula of a compound is C_2H_5 . It has a vapour density of 29. Determine the relative molecular formula mass of the compound and hence its molecular formula.

7. The percentage composition of sodium phosphate as determined by analysis, is 42.1% sodium, 18.9% phosphorus and 39% oxygen. Find the empirical formula of the compound (work to two decimal places).

8. A gaseous organic compound contains 3.6 g of carbon and 0.8 g of hydrogen. The vapour density of this compound is 22. (i) Calculate the Empirical formula. (ii) Calculate the molecular formula of the compound.

Mole Concept Review Unit 5 Mole Concept

Kevin T. Patton



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