

Name: \_\_\_\_\_ Period \_\_\_\_\_ Date \_\_\_\_\_

## Chemistry Liquids and Solids Check Off Review Answers

boiling point  
condensation point  
critical point  
critical pressure  
critical temperature  
dipole-dipole force  
dispersion force

gas  
liquid  
solid  
hydrogen bonding  
intermolecular force  
London force  
phase diagram

melting point  
freezing point  
sublimation  
deposition  
supercritical fluid  
triple point  
vapor pressure  
volatility

Use your notes and other resources to answer the following questions:

1. Discuss the Physical conversion of states of matter.  
solids → Melt → liq → evaporate → gases  
Gas → Condense → liq → freeze → solid  
Solid → Sublimate → gas; gas → deposition → solid
2. Discuss **hydrogen bonding**, **dipole-dipole**, and **London Dispersion Forces**.
  - Hydrogen bonding is a special case of dipole-dipole. Only occurs when Hydrogen is bonded to N, O, or F.
  - Molecules that Hydrogen bond also have dipole-dipole and London forces
  - London forces depend upon particle size and particle shape. Larger the particle and the larger the shape the greater the London forces. In the end, London forces come down to the number of electrons (i.e. MM).
  - All substances have London forces involved. Non-polar molecules have only London forces
3. Define dynamic equilibrium  
Dynamic Equilibrium- Two opposing processes that occur simultaneously at equal rates
4. Name three ways to increase the rate of evaporation.
  - Increasing the temperature increases the kinetic energy of the molecules in the liquid. Because they have more energy, more molecules are likely to make the transition from the liquid phase to the gas phase.
  - Lowering the pressure also increases the rate of evaporation. If you lower the external pressure, the vapor pressure necessary for evaporating - and thus the temperature necessary for evaporating - drops.
  - Increasing surface area also increases evaporation, because with a larger surface area, more molecules have a chance to make it into the vapor phase.
5. a. Why does evaporation cause cooling?  
The temperature of the water is determined by the average of the speed of all the molecules. When water evaporates at room temperature, some fast-moving, highly energetic molecules have enough energy to overcome the attractions that individual molecules have for one another and enter the gas phase. As these high-energy molecules leave the liquid phase, the molecules still in a liquid state are the slower ones (their average energy is lowered). Thus the temperature of the remaining water is lower. That's why evaporation causes cooling.
5. b. Will the remaining liquid evaporate as fast as the initial liquid?  
When the particles with the highest kinetic energy leave the substance, the average kinetic energy of the remaining particles decreases. The temperature of the substance therefore decreases. Lowering the temperature of a liquid decreases the rate of evaporation.
6. Define Vapor Pressure.  
The vapor pressure of a liquid is the equilibrium pressure of a vapor above its liquid (or solid); that is, the pressure of the vapor resulting from evaporation of a liquid above a sample of the liquid in a closed container.
7. What does the strength of vapor pressure depend on?
  - If the intermolecular force is strong, the fraction of molecules that have enough energy to escape into the vapor is small and the vapor pressure will be small.
  - Increasing the temperature increases the average kinetic energy of the liquid molecules, and hence the fraction of molecules with sufficient energy to escape from the liquid phase increases with temperature.
8. Why do liquids with low intermolecular attractive forces have high vapor pressures?

# Modern Chemistry Review Liquids To Solids Answers

**H Kauffman**



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