Nitrogen dioxide,  $NO_2$ , can be made by reacting nitrogen monoxide, NO, with oxygen,  $O_2$ .

$$2NO(g) + O_2(g) \Rightarrow 2NO_2(g)$$

A chemist mixes 0.450 mol NO with 0.450 mol  $O_2$  in a container. This container is then heated and equilibrium is reached. At equilibrium, the mixture contains 0.400 mol NO and the total pressure is 400 kPa. Calculate  $K_0$  and the units

Step 1 - Write out the equation and fill in the initial moles

	SNO	+ O₂ ≠	2NO
Initial	0.450	0.450	0
Change	-0.05	-0.025	+0.05
Equilibrium	0.400	0.425	0.05

Step 2 - Work out the partial pressures of each gas

$$p(NO) = 0.400 / 0.875 \times 400 = 182.86 \text{ kPa}$$
  
 $p(O_2) = 0.425 / 0.875 \times 400 = 194.29 \text{ kPa}$   
 $p(NO_2) = 0.05 / 0.875 \times 400 = 22.86 \text{ kPa}$ 

Step 3 – Put the values in to the  $K_p$  equation and work out the answer

$$K_p = \frac{p(NO_2)^2}{p(NO)^2 \times p(O_2)}$$
  $K_p = \frac{(22.86)^2}{(182.86)^2 \times (194.29)}$   $K_p = 8.04 \times 10^{-5}$ 

Step 4 - Calculate the units

$$K_p = \frac{(kPa)^2}{(kPa)^2}$$
  $K_p = kPa^{-1}$ 

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**B** Lingard

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understand the topic and how to apply it in the field

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