

## Assignment

Date \_\_\_\_\_ Period \_\_\_\_\_

Evaluate each expression.

$$1) 18 \div -3 - (4)((2)(4 - 1))$$

$$\text{ } -30$$

$$2) (2)(-1 - ((3)(2)) + (-3 - -2))$$

$$\text{ } 10$$

$$3) 2\frac{1}{4} \div \left( \left( (2) \left( 3\frac{3}{4} \right) \right) \left( \frac{1}{3} \right) \right)$$

$$\text{ } \frac{9}{10}$$

$$4) -3\frac{1}{5} \div -2 - \left( \frac{1}{5} \right) \left( \frac{-7}{4} \right)$$

$$\text{ } \frac{39}{20}$$

$$5) \frac{(21)(3)}{(14 - 11)^2}$$

$$\text{ } 7$$

$$6) (3) \left( 7 - \frac{(25)(3)}{15} \right)$$

$$\text{ } 6$$

Solve each inequality.

$$7) -2x - 4(x + 3) > -12 - 6x$$

$$\text{No solution. : } \begin{array}{c} \longleftarrow \text{---} | \text{---} | \text{---} | \text{---} | \text{---} | \text{---} \longrightarrow \\ 10 \quad 12 \quad 14 \quad 16 \quad 18 \quad 20 \end{array}$$

$$8) 40 - 4n \leq 2n + 4(6n - 5)$$

$$n \geq 2 : \begin{array}{c} \longleftarrow \text{---} | \text{---} | \text{---} | \bullet | \text{---} | \text{---} \longrightarrow \\ -6 \quad -2 \quad 0 \quad 2 \quad 4 \quad 6 \end{array}$$

$$9) -4 - 4r \geq -6(-r - 1)$$

$$r \leq -1 : \begin{array}{c} \longleftarrow \text{---} | \bullet | \text{---} | \text{---} | \text{---} \longrightarrow \\ -4 \quad -2 \quad 0 \quad 2 \quad 4 \end{array}$$

$$10) -2n + 16 > -6 - 2(7 + 4n)$$

$$n > -6 : \begin{array}{c} \longleftarrow \text{---} | \text{---} | \bullet | \text{---} | \text{---} \longrightarrow \\ -10 \quad -8 \quad -6 \quad -4 \quad -2 \end{array}$$

Simplify each expression.

$$11) -7(a + 10) - 3(1 + a)$$

$$\text{ } -10a - 73$$

$$12) 2(-12 + 4m) - 8m(-12m + 1)$$

$$\text{ } -24 + 96m^2$$

$$13) -10x(3 - 5x) + 12x(x + 6)$$

$$\text{ } 42x + 62x^2$$

$$14) -3(x + 9) - 5(5 + 9x)$$

$$\text{ } -48x - 52$$

# Infinite Algebra 1 Simplifying Each Expression

**R Sanford**



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web write the oxidation and reduction reactions for each redox reaction the first one is done for you  $23 \text{ Fe} + 2 \text{ Co} \rightarrow \text{Fe}_2\text{O}_3 + \text{Co}_2\text{O}_3$   
oxidation  $\text{Co} \rightarrow \text{Co}_2\text{O}_3 + 2\text{e}^-$  reduction  $\text{Fe}_2\text{O}_3 + 2\text{e}^- \rightarrow \text{Fe}$   $24 \text{ 3 Ag} + \text{Ni} \rightarrow \text{Ni}_3\text{O}_2 + 3 \text{ Ag}$  oxidation reduction  $25 \text{ Cu} + 2 \text{ Pb} \rightarrow \text{Cu}_2\text{O} + \text{Pb}_2\text{O}_3$  oxidation reduction  $26 \text{ O}_2$

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5co i2 5co2 i 2 o 5 5 c o i 2 5 c o 2

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