

Name: \_\_\_\_\_ Period: \_\_\_\_\_

### Temperature Conversions and Heat Practice

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$$K = ^\circ C + 273$$

$$^\circ F = (9/5 \times ^\circ C) + 32$$

$$^\circ C = 5/9(^{\circ}F - 32)$$

Convert the following to Fahrenheit

1)  $10^{\circ}C$  \_\_\_\_\_

2)  $30^{\circ}C$  \_\_\_\_\_

3)  $0^{\circ}C$  \_\_\_\_\_

Convert the following to Celsius

4)  $32^{\circ}F$  \_\_\_\_\_

5)  $90^{\circ}F$  \_\_\_\_\_

6)  $212^{\circ}F$  \_\_\_\_\_

Convert the following to Kelvin

7)  $0^{\circ}C$  \_\_\_\_\_

8)  $-50^{\circ}C$  \_\_\_\_\_

9)  $90^{\circ}C$  \_\_\_\_\_

10)  $-20^{\circ}C$  \_\_\_\_\_

Convert the following to Celsius

11)  $112^{\circ}K$  \_\_\_\_\_

12)  $200^{\circ}K$  \_\_\_\_\_

13)  $273^{\circ}K$  \_\_\_\_\_

14)  $350^{\circ}K$  \_\_\_\_\_

Use the information below to complete the following problems

Heat = mass x specific heat x change in temperature

$\Delta t$  = final temperature – initial temperature

$$Q = m c \Delta t$$

$$\Delta t = t_f - t_i$$

| Substances     | (j/kg•K) | Substance | (j/kg•K) | Substance | (j/kg•K) |
|----------------|----------|-----------|----------|-----------|----------|
| Water (liquid) | 4186     | Copper    | 385      | Iron      | 449      |
| Steam          | 1870     | Gold      | 129      | Aluminum  | 897      |

15) How much energy is needed to heat 55 Kg of water from 285 K to 360 K?

16) How much energy is released when 55 Kg of steam cools from 558 K to 390 K?

17) If 5785 J of energy is used to heat 48 kg of Iron what is it's change in temperature?

18) An 75 kg block of metal absorbs 721875 J of heat as the temperature increases from 375 K to 400 K.  
What is the specific heat of the metal?

a. Which metal is it likely to be?

19) How much energy is released if 278 kG of Aluminum cools from 698 K to 298 K?

20) What is the specific heat of a 3.78 kg object that absorbs 678 J as the temperature increases by 4.25 K?

# Het And Its Measurement Answer Key

**WJ Hussar**



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