



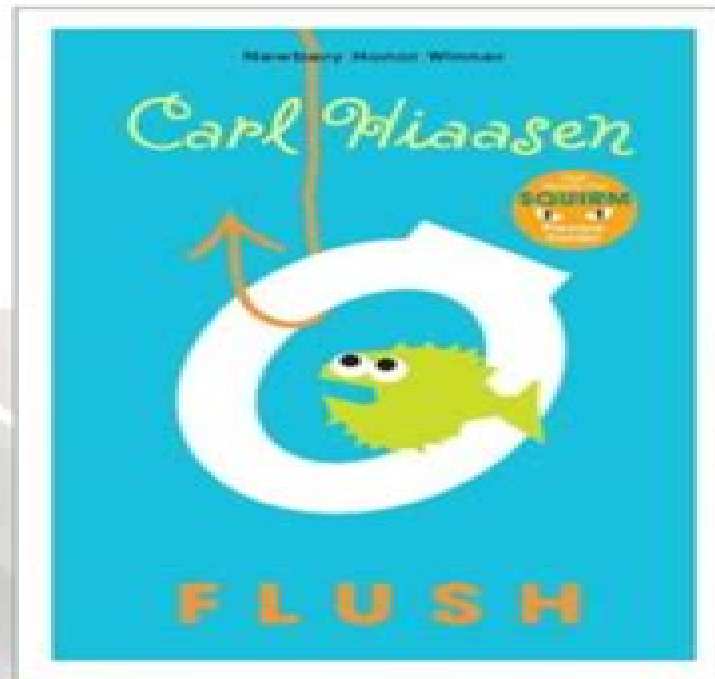
TEACHER GUIDE

GRADES 6-8

COMPREHENSIVE CURRICULUM BASED LESSON PLANS

Flush

Carl Hiaasen



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web we utilize the equation that relate amount remaining initial mass and number of half lives $n = \frac{\ln \left(\frac{N_0}{N} \right)}{\ln 2}$ $N = N_0 \left(\frac{1}{2} \right)^n$ $125 \text{ g} \times \left(\frac{1}{2} \right)^5$ what is the half life of an isotope that is 80 remained after 16 days solution remaining 80 100 therefore $n = \frac{\ln \left(\frac{100}{80} \right)}{\ln 2}$ now using the half life equation in example 1 we have

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