

25. Given the reaction:

$$4HCl_{(g)} + O_{2(g)} \rightleftharpoons 2H_2O_{(g)} + 2Cl_{2(g)} \Delta H = -113 \text{ kJ}$$

How will the value of the equilibrium constant Keq at 550°C compare with it's value at

45000 Keg at 550°c will have a lower value.

the reaction to shift LEFT. This will decrease the value of keg.

26. The following system is at equilibrium, in a closed container:

- a) How is the amount of N2 in the container affected if the volume of the container is doubled? Increased (decreasing Pwill cause a shift RIGHT producing more N2)
- b) How is the rate of the forward reaction affected if more water vapor is introduced into the container? Increase (First inc. [H20] will speed up rev.rx, forming more leactants. Then forward rx, will speed up.
- c) How is the amount of O2 in the container affected if a catalyst is added?

No change (A catalyst speeds up forward frev. rx. equally)

27. At a certain temperature, Keq for the reaction:

If the equilibrium concentration of C_2H_2 is 0.40 moles/L, what is the equilibrium concentration of C_6H_6 ?

Answer 0.32 M

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