

**PUNJAB PUBLIC SERVICE COMMISSION
COMBINED COMPETITIVE EXAMINATION
FOR RECRUITMENT TO THE POSTS OF
PROVINCIAL MANAGEMENT SERVICE, ETC.**

CHEMISTRY (OPTIONAL) PAPER-I

TIME ALLOWED: 3 HOURS

MAXIMUM MARKS: 100

Note: Attempt any FIVE questions only. Write answers in English.

- Q.1:** (a) Using a diagram, explain the variation of molecular energy with internuclear separation. (10)
- (b) What is the primary assumption of the VSEPR model? Explain the difference between square-pyramidal and trigonal bi-pyramidal arrangement in terms of VSEPR. Predict the shape of an SeF_4 molecule. (10)
- Q.2:** (a) What are the mineral sources for copper? Describe in detail the information which can be obtained for 3d metals using a FROST diagram. (10)
- (b) (i) On the basis of trends in the 3d series, suggest possible M^{2+} aqua ions for use as reducing agents, and write a balanced equation for the reaction of one of these ions with O_2 in acidic solution. (5)
- (ii) Comment on the 2+ oxidation states of 4d and 5d elements. (5)
- Q.3:** (a) Explain with the help of a flow sheet diagram, the electrolytic process of caustic soda manufacturing. (10)
- (b) Give a detailed comparison between dry and wet processing of cement manufacturing. (10)
- Q.4:** (a) Discuss the physical origin of quantization energy for a particle confined to moving inside a one dimensional box or on a ring. (10)
- (b) Use EHT to calculate the energies of the three molecular orbitals in an H_3^+ molecule ion treated as an equilateral triangle of side $R = 3.25a_0$ (170 pm). (10)
- Q.5:** (a) Calculate the work done when 50g of iron reacts with HCl to produce hydrogen gas in: (5)
- (i) a closed vessel of fixed volume, (ii) an open beaker at $25^\circ C$.
- (b) The molar constant pressure heat capacity of a certain solid at 10K is $0.43 J K^{-1} mol^{-1}$. What is its molar entropy at that temperature? (10)
- Q.6:** (a) Suggest a calibration procedure for a Na^+ selective electrode, in which the values of E_{ap} , β and k_{xy} are measured experimentally. Assume that K^+ is the only interfering species. (10)
- (b) Distinguish between galvanic and electrolytic cells. (10)
- Q.7:** (a) Define Nuclear Fission and Nuclear Fusion processes. Discuss their comparative features. (10)
- (b) Outline a method for preparation of CAN pellets. Distinguish between calcium superphosphate and Triple Superphosphate. (10)

June 2012 66ch01 Chemistry Past Paper

RJ Alexander



June 2012 66ch01 Chemistry Past Paper:

The book delves into June 2012 66ch01 Chemistry Past Paper. June 2012 66ch01 Chemistry Past Paper is a vital topic that must be grasped by everyone, from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into June 2012 66ch01 Chemistry Past Paper, encompassing both the fundamentals and more intricate discussions.

1. The book is structured into several chapters, namely:

- Chapter 1: Introduction to June 2012 66ch01 Chemistry Past Paper
- Chapter 2: Essential Elements of June 2012 66ch01 Chemistry Past Paper
- Chapter 3: June 2012 66ch01 Chemistry Past Paper in Everyday Life
- Chapter 4: June 2012 66ch01 Chemistry Past Paper in Specific Contexts
- Chapter 5: Conclusion

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