

Graphing Periodic Trends Activity

Name: _____ Period: _____

Use the information in these tables to complete the graph as described below

Symbol	Atomic Radius (Picometers)	First Ionization energy (kilojoules/mole)	Electronegativity (4-point scale)
H	31	1312	2.1
Li	128	520	1.0
Na	166	496	0.9
K	203	410	0.8
Rb	220	403	0.8
Cs	244	376	0.7

Symbol	Atomic Radius (Picometers)	First Ionization energy (kilojoules/mole)	Electronegativity (4-point scale)
Na	166	496	0.9
Mg	141	738	1.2
Al	121	578	1.5
Si	111	787	1.8
P	107	1012	2.1
S	105	1000	2.5
Cl	102	1251	3.0
Ar	106	1521	---

1. Using colored pencils or pen, list under the symbol (in this order!) the Atomic Radius, First Ionization Energy, and Electronegativity.
2. Use a different color for each property. Example: Write all of the Atomic radius values in red, all of the First ionization energies in green, and all of the Electronegativities in blue. Of course, you can pick any colors that are available, as long as you are consistent.
3. Observe the trends in each property as you go down the Alkali metal group, and as you go across Period 3.
4. Write out each of the statements written on the opposite side, completing each statement with the observed trend (increase or decrease)

H	Across a period (→) atomic radius tends to _____. Across a period (→) first ionization energy tends to _____. Across a period (→) electronegativity tends to _____.						
Li							
Na 166 496 0.9	Mg	Al	Si	P	S	Cl	Ar
K	Down a Group (↓) atomic radius tends to _____. Down a Group (↓) first ionization energy tends to _____. Down a Group (↓) electronegativity tends to _____.						
Rb							
Cs							

Graphing Periodic Trends Worksheet Answerkey

Daniel F McAuley



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Table of Contents Graphing Periodic Trends Worksheet Answerkey

1. Understanding the eBook Graphing Periodic Trends Worksheet Answerkey
 - The Rise of Digital Reading Graphing Periodic Trends Worksheet Answerkey
 - Advantages of eBooks Over Traditional Books
2. Identifying Graphing Periodic Trends Worksheet Answerkey
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Graphing Periodic Trends Worksheet Answerkey
 - User-Friendly Interface
4. Exploring eBook Recommendations from Graphing Periodic Trends Worksheet Answerkey
 - Personalized Recommendations
 - Graphing Periodic Trends Worksheet Answerkey User Reviews and Ratings
 - Graphing Periodic Trends Worksheet Answerkey and Bestseller Lists
5. Accessing Graphing Periodic Trends Worksheet Answerkey Free and Paid eBooks
 - Graphing Periodic Trends Worksheet Answerkey Public Domain eBooks
 - Graphing Periodic Trends Worksheet Answerkey eBook Subscription Services
 - Graphing Periodic Trends Worksheet Answerkey Budget-Friendly Options

6. Navigating Graphing Periodic Trends Worksheet Answerkey eBook Formats
 - ePub, PDF, MOBI, and More
 - Graphing Periodic Trends Worksheet Answerkey Compatibility with Devices
 - Graphing Periodic Trends Worksheet Answerkey Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Graphing Periodic Trends Worksheet Answerkey
 - Highlighting and Note-Taking Graphing Periodic Trends Worksheet Answerkey
 - Interactive Elements Graphing Periodic Trends Worksheet Answerkey
8. Staying Engaged with Graphing Periodic Trends Worksheet Answerkey
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Graphing Periodic Trends Worksheet Answerkey
9. Balancing eBooks and Physical Books Graphing Periodic Trends Worksheet Answerkey
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Graphing Periodic Trends Worksheet Answerkey
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Graphing Periodic Trends Worksheet Answerkey
 - Setting Reading Goals Graphing Periodic Trends Worksheet Answerkey
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Graphing Periodic Trends Worksheet Answerkey
 - Fact-Checking eBook Content of Graphing Periodic Trends Worksheet Answerkey
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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