

GRAPHING PERIODIC TRENDS

Standard: Students know how to use the periodic table to identify trends in ionization energy, electronegativity, and the relative sizes of ions and atoms.

PRE-LAB DISCUSSION:

The Periodic Table is arranged according to the Periodic Law. The Periodic Law states that when elements are arranged in order of increasing atomic number, their physical and chemical properties show a periodic pattern. Students can discover these patterns by examining the changes in properties of elements on the Periodic Table. The properties that will be examined in this lesson are: atomic radius AND first ionization energy.

PURPOSE:

To understand periodic trends relating to atomic radius, ionization energy, and electronegativity.

PROCEDURE:

Use the information in these tables to complete the graph as described below.

RESULTS:

Symbol	Atomic Radius (Picometers)	First Ionization (Kilojoules/mole)	Electronegativity (4-point scale)
H	37	1312	2.1
Li	128	520	1.0
Na	190	496	0.9
K	203	419	0.8
Rb	220	403	0.8
Cs	244	376	0.7

Symbol	Atomic Radius (Picometers)	First Ionization (Kilojoules/mole)	Electronegativity (4-point scale)
He	31	498	0.0
Mg	141	738	1.2
Al	121	578	1.5
Si	111	787	1.8
P	107	1012	2.1
S	106	1000	2.5
Cl	102	1251	3.0
Ar	100	1521	—

1. Define each of these terms: Atomic radius, First Ionization Energy, and Electronegativity.
2. Using colored-pencil or pen, list under the symbol (in this order) the Atomic Radius, First Ionization Energy, and Electronegativity. Use a different color for each property. Example: Write all of the Atomic radius values in red, all of the First Ionization energies in green, and all of the Electronegativities in blue. Of course, you can pick any colors that are available, as long as you are consistent.
3. Observe the trends in each property as you go down the Alkali metal group, and as you go across Period 3.
4. Write out each of the statements written on the opposite side, completing each statement with the observed trend (increase or decrease).
5. There is NO ABSTRACT to write for this activity :)

Graphing Periodic Relationships Lab Answer Key

Bowker Editorial Staff



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Table of Contents Graphing Periodic Relationships Lab Answer Key

1. Understanding the eBook Graphing Periodic Relationships Lab Answer Key
 - The Rise of Digital Reading Graphing Periodic Relationships Lab Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Graphing Periodic Relationships Lab Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Graphing Periodic Relationships Lab Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Graphing Periodic Relationships Lab Answer Key
 - Personalized Recommendations

- Graphing Periodic Relationships Lab Answer Key User Reviews and Ratings
- Graphing Periodic Relationships Lab Answer Key and Bestseller Lists
- 5. Accessing Graphing Periodic Relationships Lab Answer Key Free and Paid eBooks
 - Graphing Periodic Relationships Lab Answer Key Public Domain eBooks
 - Graphing Periodic Relationships Lab Answer Key eBook Subscription Services
 - Graphing Periodic Relationships Lab Answer Key Budget-Friendly Options
- 6. Navigating Graphing Periodic Relationships Lab Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Graphing Periodic Relationships Lab Answer Key Compatibility with Devices
 - Graphing Periodic Relationships Lab Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Graphing Periodic Relationships Lab Answer Key
 - Highlighting and Note-Taking Graphing Periodic Relationships Lab Answer Key
 - Interactive Elements Graphing Periodic Relationships Lab Answer Key
- 8. Staying Engaged with Graphing Periodic Relationships Lab Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Graphing Periodic Relationships Lab Answer Key
- 9. Balancing eBooks and Physical Books Graphing Periodic Relationships Lab Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Graphing Periodic Relationships Lab Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Graphing Periodic Relationships Lab Answer Key
 - Setting Reading Goals Graphing Periodic Relationships Lab Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Graphing Periodic Relationships Lab Answer Key
 - Fact-Checking eBook Content of Graphing Periodic Relationships Lab Answer Key

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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