

# GRAM FORMULA MASS

Name Lia Kerner

Determine the gram formula mass (the mass of one mole) of each compound below.

$\frac{1}{2} \times 16 = 8$   
 $3 \times 8 = 24$   
 $18$

1.  $\text{KMnO}_4$  158 g  
 $39 + 55 + 4(16) = 158$

2.  $\text{KCl}$  74 g  
 $39 + 35 = 74$

3.  $\text{Na}_2\text{SO}_4$  142 g  
 $2(23) + 32 + 4(16) = 142$

4.  $\text{Ca(NO}_3)_2$  164 g  
 $40 + 2(14 + 3 \times 16) = 164$

5.  $\text{Al}_2(\text{SO}_4)_3$  342 g  
 $2(27) + 3(32 + 4 \times 16) = 342$

6.  $(\text{NH}_4)_3\text{PO}_4$  149 g  
 $3(14 + 4 \times 1) + 31 + 4 \times 16 = 149$

7.  $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$  250 g  
 $64 + 96 + 5(2 + 16) = 250$

8.  $\text{Mg}_3(\text{PO}_4)_2$  264 g  
 $3(24) + (96) \times 2 = 264$

9.  $7\text{H}_2(\text{C}_2\text{H}_3\text{O}_2) \cdot 2\text{H}_2\text{O}$  219 g

# Gram Formula Mass Answers

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