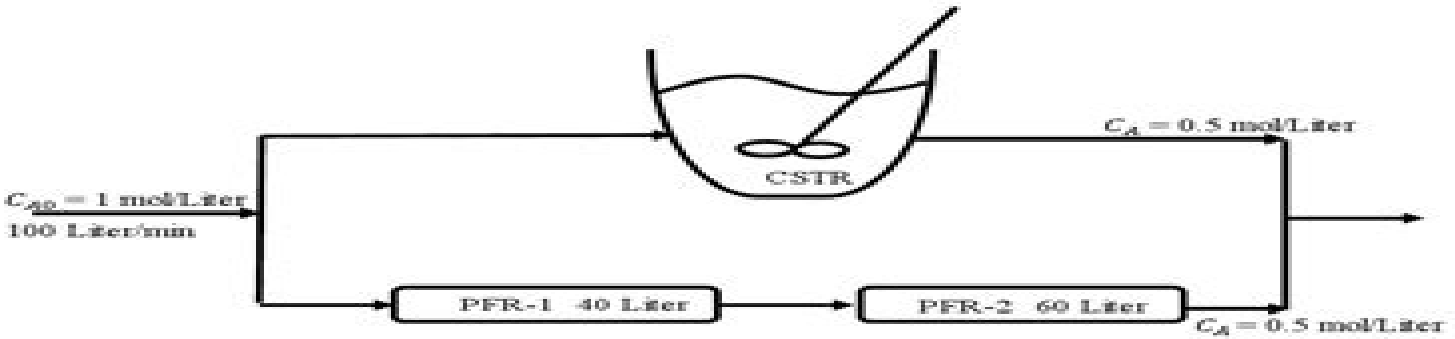
Q.46 A first-order irreversible liquid phase reaction $A \rightarrow B$ ($k = 0.1\ \text{min}^{-1}$) is carried out under isothermal, steady state conditions in the following reactor arrangement comprising an ideal CSTR (Continuous-Stirred Tank Reactor) and two ideal PFRs (Plug Flow Reactors). From the information in the figure, the volume of the CSTR (in Liters) is _____ (rounded off to the nearest integer).



Q.47 The elementary liquid-phase irreversible reactions

$$A \xrightarrow{k_1 = 0.4\ \text{min}^{-1}} B \xrightarrow{k_2 = 0.1\ \text{min}^{-1}} C$$

take place in an isothermal ideal CSTR (Continuous-Stirred Tank Reactor). Pure A is fed to the reactor at a concentration of 2 mol/Liter. For the residence time that maximizes the exit concentration of B, the percentage yield of B, defined as $\left(\frac{\text{net formation rate of B}}{\text{consumption rate of A}} \times 100\right)$, is _____ (rounded off to the nearest integer).

Q.48 The elementary irreversible gas-phase reaction A → B + C is carried out adiabatically in an ideal CSTR (Continuous-Stirred Tank Reactor) operating at 10 atm. Pure A enters the CSTR at a flow rate of 10 mol/s and a temperature of 450 K. Assume A, B and C to be ideal gases. The specific heat capacity at constant pressure (C_{pi}) and heat of formation (H_i°), of component i (i = A, B, C), are:

$C_{pA} = 30$ J/(mol K) $\quad C_{pB} = 10$ J/(mol K) $\quad C_{pC} = 20$ J/(mol K)

$H_A^\circ = -90$ kJ/mol $\quad H_B^\circ = -54$ kJ/mol $\quad H_C^\circ = -45$ kJ/mol

The reaction rate constant $k\ (\text{per second}) = 0.133\exp\left\{\frac{E}{R}\left(\frac{1}{450} - \frac{1}{T}\right)\right\}$, where $E = 31.4$ kJ/mol and universal gas constant $R = 0.082$ L atm/(mol K) $= 8.314$ J/(mol K). The shaft work may be neglected in the analysis, and specific heat capacities do not vary with temperature. All heats of formation are referenced to 273 K. The reactor volume (in Liters) for 75% conversion is _____ (rounded off to the nearest integer).

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