Guizizz	NAME:	
Stoichiometry Practice 10 Questions	CLASS:	
	DATE :	

1	Balance the following reaction:						
	Mg(s) + HCl(aq)> MgC	:I ₂ (aq) +	$H_2(g)$				

- 1.2.1.2 1,2,2,2
- 2,1,1,2 1,2,1,1
- 2 Using the balanced equation in question 1, $Mg(s) + HCl(aq) \rightarrow MgCl_2(aq) + H_2(g)$ How many grams of HCI are consumed by the reaction of 2.50 moles of magnesium?
- A. 218 g 7.29
- 5.00 g 182 g
 - Using the balanced equation in question 1, $Mg(s) + HCl(aq) \rightarrow MgCl_2(aq) + H_2(g)$ What is the mass in grams of H₂ gas when 4.0 moles of HCI is added to the reaction?
- A $0.99\,\mathrm{g}$ 4.0 g C 2.0 g 6.0 g
- Using the balanced equation in question 1, $Mg(s) + HCl(aq) \rightarrow MgCl_2(aq) + H_2(g)$ How many moles of HCI are consumed in the production of 7.5 moles of MgCl₂?

7.5 moles:

- 15 moles 3.8 moles C (D)
- Balance the following reaction: $CaC_2(s) + H_2O(l) --> C_2H_2(g) + Ca(OH)_2(aq)$

23 moles

- 2, 1, 1, 1 2,1,2,1
- 1,2,2,2 1,2,1,1

Grade 11 Physics Exam Questions Stoichiometry

A Gutmann

Grade 11	l Physics	Exam (Questions 9	Stoichiome	try:	

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