

How Substances Dissolve

KEY IDEAS

As you read this section, keep these questions in mind:

- Why is water called the universal solvent?
- How do substances dissolve?

Why Do Substances Dissolve in Certain Solvents?

Water is often called a universal solvent because many substances dissolve in it. However, no one substance can dissolve every solute. A general rule in chemistry is that “like dissolves like.” This rule means that a solvent will dissolve substances that have similar molecular structures. 

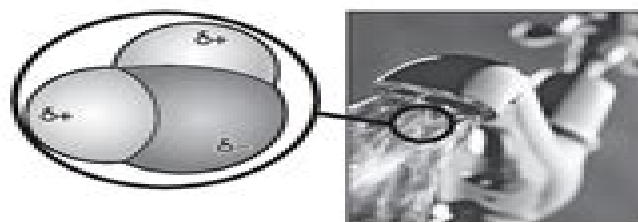
A solvent must be able to attract solute particles and pull them away from one another. Thus, solvent particles must attract solute particles more strongly than the solute particles attract one another.

POLAR MOLECULES

Recall that atoms bonded together do not all share electrons equally. For example, a water molecule is made up of two hydrogen atoms bonded to one oxygen atom. The oxygen atom attracts electrons more strongly than the hydrogen atoms do. Thus, the electrons spend more time near the oxygen atom. This gives each hydrogen atom a partial positive charge (δ^+) and the oxygen atom a partial negative charge (δ^-). 

Molecules that have partially charged positive and negative areas are **polar**. Because water molecules are polar, water can dissolve many other polar substances. Water can also dissolve ionic compounds such as table salt.

Water is a polar molecule because the oxygen atom strongly attracts electrons. Because the electrons spend more time near the oxygen atom, the oxygen atom has a partial negative charge.



READING TOOLBOX

Summarize Read each page of the section silently to yourself. With a partner, take turns summarizing each page.

READING CHECK

1. Identify Why is water called a universal solvent?

because
many substances dissolve
in it

READING CHECK

2. Explain Why does a water molecule have partial charges?

The oxygen atom attracts
electrons more strongly
than the hydrogen atoms
do. Thus, the electrons
spend more time near the
oxygen atom.

How Substances Dissolve Concept Review Answers

RC Schank



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Table of Contents How Substances Dissolve Concept Review Answers

1. Understanding the eBook How Substances Dissolve Concept Review Answers
 - The Rise of Digital Reading How Substances Dissolve Concept Review Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying How Substances Dissolve Concept Review Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an How Substances Dissolve Concept Review Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from How Substances Dissolve Concept Review Answers
 - Personalized Recommendations
 - How Substances Dissolve Concept Review Answers User Reviews and Ratings
 - How Substances Dissolve Concept Review Answers and Bestseller Lists
5. Accessing How Substances Dissolve Concept Review Answers Free and Paid eBooks
 - How Substances Dissolve Concept Review Answers Public Domain eBooks
 - How Substances Dissolve Concept Review Answers eBook Subscription Services

- How Substances Dissolve Concept Review Answers Budget-Friendly Options
- 6. Navigating How Substances Dissolve Concept Review Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - How Substances Dissolve Concept Review Answers Compatibility with Devices
 - How Substances Dissolve Concept Review Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of How Substances Dissolve Concept Review Answers
 - Highlighting and Note-Taking How Substances Dissolve Concept Review Answers
 - Interactive Elements How Substances Dissolve Concept Review Answers
- 8. Staying Engaged with How Substances Dissolve Concept Review Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers How Substances Dissolve Concept Review Answers
- 9. Balancing eBooks and Physical Books How Substances Dissolve Concept Review Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection How Substances Dissolve Concept Review Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine How Substances Dissolve Concept Review Answers
 - Setting Reading Goals How Substances Dissolve Concept Review Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of How Substances Dissolve Concept Review Answers
 - Fact-Checking eBook Content of How Substances Dissolve Concept Review Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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