

Hewlett-Packard Professional Certification Assessment

Examination 1017 has six to sixteen following problems. Use all appropriate and proper scientific units. Use scientific notation for answers requiring scientific notation in the answer blank.

1. Find the standard enthalpy change for the reaction $\text{C}_{10}\text{H}_8 + 12\text{O}_2 \rightarrow 10\text{CO}_2 + 8\text{H}_2\text{O}$
 of carbon (s) $\Delta H_{\text{f}}^\circ = 0 \text{ kJ/mol}$ $\Delta H_{\text{f}}^\circ = -393.5 \text{ kJ/mol}$
 and C_{10}H_8 $\Delta H_{\text{f}}^\circ = 0 \text{ kJ/mol}$ $\Delta H_{\text{f}}^\circ = -125.6 \text{ kJ/mol}$

a. Correct $\Delta H_{\text{rxn}}^\circ = 125.6 \text{ kJ}$ $\Delta H_{\text{rxn}}^\circ = -125.6 \text{ kJ}$ $\Delta H_{\text{rxn}}^\circ = -1256 \text{ kJ}$
 b. Correct $\Delta H_{\text{rxn}}^\circ = 1256 \text{ kJ}$ $\Delta H_{\text{rxn}}^\circ = -1256 \text{ kJ}$ $\Delta H_{\text{rxn}}^\circ = -125.6 \text{ kJ}$

c. Correct $\Delta H_{\text{rxn}}^\circ = 125.6 \text{ kJ}$ $\Delta H_{\text{rxn}}^\circ = -125.6 \text{ kJ}$ $\Delta H_{\text{rxn}}^\circ = -1256 \text{ kJ}$

2. The standard enthalpy change for the formation of aluminum oxide and aluminum hydroxide are -1676 kJ/mol $\Delta H_{\text{f}}^\circ = -1676 \text{ kJ/mol}$ $\Delta H_{\text{f}}^\circ = -1676 \text{ kJ/mol}$ $\Delta H_{\text{f}}^\circ = -1676 \text{ kJ/mol}$
 $\Delta H_{\text{f}}^\circ = -1676 \text{ kJ/mol}$ $\Delta H_{\text{f}}^\circ = -1676 \text{ kJ/mol}$ $\Delta H_{\text{f}}^\circ = -1676 \text{ kJ/mol}$ $\Delta H_{\text{f}}^\circ = -1676 \text{ kJ/mol}$
 Calculate the standard enthalpy change for the reaction $2\text{Al} + 3\text{H}_2\text{O} \rightarrow \text{Al}_2\text{O}_3 + 3\text{H}_2$

a. Correct $\Delta H_{\text{rxn}}^\circ = 1676 \text{ kJ}$ $\Delta H_{\text{rxn}}^\circ = -1676 \text{ kJ}$ $\Delta H_{\text{rxn}}^\circ = -1676 \text{ kJ}$
 b. Correct $\Delta H_{\text{rxn}}^\circ = 1676 \text{ kJ}$ $\Delta H_{\text{rxn}}^\circ = -1676 \text{ kJ}$ $\Delta H_{\text{rxn}}^\circ = -1676 \text{ kJ}$

c. Correct $\Delta H_{\text{rxn}}^\circ = 1676 \text{ kJ}$ $\Delta H_{\text{rxn}}^\circ = -1676 \text{ kJ}$ $\Delta H_{\text{rxn}}^\circ = -1676 \text{ kJ}$
 d. Correct $\Delta H_{\text{rxn}}^\circ = 1676 \text{ kJ}$ $\Delta H_{\text{rxn}}^\circ = -1676 \text{ kJ}$ $\Delta H_{\text{rxn}}^\circ = -1676 \text{ kJ}$

e. Correct $\Delta H_{\text{rxn}}^\circ = 1676 \text{ kJ}$ $\Delta H_{\text{rxn}}^\circ = -1676 \text{ kJ}$ $\Delta H_{\text{rxn}}^\circ = -1676 \text{ kJ}$

3. Find the standard enthalpy change for the combustion reaction of carbon monoxide and hydrogen, which would give the equation $\text{CO} + \text{H}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$ $\Delta H_{\text{f}}^\circ = 0 \text{ kJ/mol}$

Calculate the standard enthalpy change for this reaction from the following standard enthalpies:

$\Delta H_{\text{f}}^\circ = 0 \text{ kJ/mol}$
 $\Delta H_{\text{f}}^\circ = 0 \text{ kJ/mol}$ $\Delta H_{\text{f}}^\circ = 0 \text{ kJ/mol}$ $\Delta H_{\text{f}}^\circ = 0 \text{ kJ/mol}$
 $\Delta H_{\text{f}}^\circ = 0 \text{ kJ/mol}$ $\Delta H_{\text{f}}^\circ = 0 \text{ kJ/mol}$ $\Delta H_{\text{f}}^\circ = 0 \text{ kJ/mol}$

a. Correct $\Delta H_{\text{rxn}}^\circ = 0 \text{ kJ}$ $\Delta H_{\text{rxn}}^\circ = 0 \text{ kJ}$ $\Delta H_{\text{rxn}}^\circ = 0 \text{ kJ}$
 b. Correct $\Delta H_{\text{rxn}}^\circ = 0 \text{ kJ}$ $\Delta H_{\text{rxn}}^\circ = 0 \text{ kJ}$ $\Delta H_{\text{rxn}}^\circ = 0 \text{ kJ}$
 c. Correct $\Delta H_{\text{rxn}}^\circ = 0 \text{ kJ}$ $\Delta H_{\text{rxn}}^\circ = 0 \text{ kJ}$ $\Delta H_{\text{rxn}}^\circ = 0 \text{ kJ}$

d. Correct $\Delta H_{\text{rxn}}^\circ = 0 \text{ kJ}$ $\Delta H_{\text{rxn}}^\circ = 0 \text{ kJ}$ $\Delta H_{\text{rxn}}^\circ = 0 \text{ kJ}$
 e. Correct $\Delta H_{\text{rxn}}^\circ = 0 \text{ kJ}$ $\Delta H_{\text{rxn}}^\circ = 0 \text{ kJ}$ $\Delta H_{\text{rxn}}^\circ = 0 \text{ kJ}$

f. Correct $\Delta H_{\text{rxn}}^\circ = 0 \text{ kJ}$ $\Delta H_{\text{rxn}}^\circ = 0 \text{ kJ}$ $\Delta H_{\text{rxn}}^\circ = 0 \text{ kJ}$

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