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Section \_\_\_\_\_ Team \_\_\_\_\_  
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## Pre-Lab Study Questions | 14

1. Why does an oil-and-vinegar salad dressing have two separate layers?

Salad dressing forms layers because oil is not soluble in water. This causes a heterogeneous mixture, which is seen as two separate layers.

2. What is meant by the mass percent (m/m) concentration of a solution?

It describes the grams of solute in the grams of solution multiplied by 100.

3. What is molarity?

It is a concentration unit. It is the moles of solute in the liters of solution.

4. Why are some electrolytes considered strong, whereas others are considered weak?

Strong electrolytes, such as strong acids, form many ions in solution.

Weak electrolytes, such as weak acids, form few ions in solution.

5. A solution is prepared with 3.26 g KCl and water to make 25.0 mL of KCl.

- a. What is the % (m/v) of the KCl solution?

$$\%(\text{m}/\text{v})\text{KCl} = \frac{3.26\text{g KCl}}{25.0\text{mL sol'n}} (100) = 13.04 \approx 13.0\%$$

- b. What is the molarity (M) of the KCl solution?

$$M_{\text{KCl}} = \frac{3.26\text{g KCl} \left( \frac{1\text{ mole KCl}}{74.6\text{g KCl}} \right)}{25.0\text{ mL sol'n} \left( \frac{1\text{ L}}{1000\text{ mL}} \right)} = 1.747989 \approx 1.75\text{ M KCl}$$

# Epcc Laboratory Experiments Lab Manual Answers

**J Dewey**



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