- Which event must. always occur for a chemical reaction to take place?
 - A. formation of a precipitate
 - B. formation of a gas-
 - C. effective collisions between reacting particles
 - D. addition of a catalyst to the reaction system
- In most aqueous reactions as temperature increases, the effectiveness of collisions between reacting particles
 - A. decreases

C. remains the same

- B. increases
- A 5.0-gram sample of zinc and a 50,-millil iter sample of hydrochloric acid are used in a chemical reaction. Which combination of these samples has the fastest reaction rate?
 - A. a rinc strip and 1.0 M HCl(aq):
 - B. a zinc strip and 3.0 M HCl(aq)
 - C. zinc powder and I.0 M HCl(aq)
 - D. zinc powder and 3.0 M HCl(aq)
- 4. A 1.0-gram piece of zinc reacts with 5 millil iters of HCl(aq). Which of these conditions of concentration and temperat ure would produce the greatest rate of reaction?
 - A. 1.0 M HCl(aq) at 20.°C
- C. 2.0 M HCl(aq) at 20.5C
- 1.0 M HCl(aq) at 40.°C
- D: 2.0 M HCl(aq) at 40.°C
- At STP, which 4.0-gram a inc sample will react fastest with dilute hydrochloric acid?
 - A. hump

C. powdered

D. bar

- D. sheet metal
- Given the balanced equation representing a reaction:

$$Zn(s) = 2HCl(aq) \rightarrow H_s(g) + ZnCl_s(aq)$$

Which set of reaction conditions produces H 2(g) at the fastest rate?

- A. a 1.0-g lump of Zn(s) in 50. mL of 0.5 M HC1(aq) at 20.°C
- B. a 1.0-g lump of Zn(s) in 50. ml. of 0.5 M HC1(aq) at 30.°C
- C. 1.0 g of powdered Zn(x) i n 50. mL of 1.0 M HC1(aq) at 20.5C
- D. 1.0 g of powdered Zn(x) i n 50. mL of 1.0 M HC1(aq) at 30. °C
- For a given reaction, adding a catalyst increases the rate of the reaction by
 - providing an alternate reaction pathway that has a higher activation energy
 - B. providing an alternate reaction pathway that has a lower activation energy
 - using the same reaction pathway and increasing the activation energy
 - using the same reaction pathway and decreasing the activation energy

8. Given the balanced equation represent ing a reaction:

$$CH_{\nu}(g) + 2O_{\nu}(g) \rightarrow 2H_{\nu}O(g) + CO_{\nu}(g) + heat$$

Which statement is true about energy in this reaction?

- A. The reaction is exothermic because it releases heat.
- B. The reaction is exothermic because it absorbs heat.
- C. The reaction is endothermic because it releases heat.
- D. The reaction is endothermic because it absorbs heat.
- 9. Which balanced equation represents an endothermic reaction?
 - A. $C(s) + O_{\gamma}(g) \rightarrow CO_{\gamma}(g)$
 - B. $CH_{s}(g) + 2O_{s}(g) \rightarrow CO_{s}(g) + 2H_{s}O(n)$
 - C. $N_1(g) + 3H_2(g) \rightarrow 2NH_2(g)$
 - D. $N_j(g) + O_j(g) \rightarrow 2NO(g)$
- 10. Given the balanced equation:

$$KNO_3(s) + 34.89 \text{ kJ} \xrightarrow{H_2O} K^*(aq) + NO_3^*(aq)$$

Which statement best describes this process?

- It is endothermic and entropy increases.
- It is endothermic and entropy decreases.
- It is exothermic and entropy increases.
- D. It is exothermic and entropy decreases.
- 11. Which change is exothermic?
 - A. freezing of water
- C. vaporization of ethanol
- B. melting of iron
- D. sublimation of iodine
- 12. Which reaction releases the greatest amount of energy per 2 moles of product?
 - A. 2CO(g) + O₂(g) → 2CO₂(g)
 - B. 4Al(s) + 3O₃(g) -> 2Al₂O₃(s)
 - $C_- = 2H_2(g) \pm O_2(g) \longrightarrow 2H_2O(g)$
 - D. $N_2(g) + 3H_2(g) \rightarrow 2NH_2(g)$
- Given the balanced equation representing a reaction at 101.3 kPa and 298 K:

$$N_2(g) + 3H_2(g) \rightarrow 2NH_2(g) + 91.8 \text{ kJ}$$

Which statement is true about this reaction?

- A. It is exothermic and AII equals -91.8 kJ.
- B. It is exothermic and AH equals +91.8 kJ.
- C. It is endothermic and AH equals –91.8 kJ.
- D. It is endothermic and ΔH equals +91.8 kJ.
- 14. Which expression represents the AH for a chemical reaction in terms of the potential ener gy, PE, of its products and reactants?
 - A. PE of products + PE of reactants
 - B. PE of products PE of reactants
 - C. PE of products × PE of reactants
 - D. PE of products + PE of reactants

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