

# CHEMISTRY



## Periodic Table of Elements

1 H 1.008																	18 Ar 39.948
2 He 4.0026																	36 Kr 83.80
3 Li 6.941	4 Be 9.0122															54 Xe 131.29	86 Rn 222
5 Na 22.990	6 Mg 24.305															84 Po 209	118 Og 294
7 K 39.098	8 Ca 40.078	9 Sc 44.956	10 Ti 47.88	11 V 50.942	12 Cr 52.00	13 Mn 54.938	14 Fe 55.845	15 Co 58.933	16 Ni 58.69	17 Cu 63.546	18 Zn 65.38	19 Ga 69.723	20 Ge 72.64	21 As 74.922	22 Se 78.96	23 Br 79.904	24 Kr 83.80
19 Rb 85.468	20 Sr 87.62	21 Y 88.906	22 Zr 91.224	23 Nb 92.906	24 Mo 95.94	25 Tc 98	26 Ru 101.07	27 Rh 101.07	28 Pd 106.90	29 Ag 107.868	30 Cd 112.411	31 In 114.818	32 Sn 118.710	33 Sb 121.757	34 Te 127.6	35 I 126.905	36 Xe 131.29
37 Cs 132.905	38 Ba 137.327	39 La 138.905	40 Hf 178.49	41 Ta 180.948	42 W 183.84	43 Re 186.207	44 Os 190.23	45 Ir 192.222	46 Pt 195.084	47 Au 196.967	48 Hg 200.59	49 Tl 204.383	50 Pb 207.2	51 Bi 208.980	52 Po 209	53 At 210	54 Rn 222
55 Fr 223	56 Ra 226	57 Ac 227	58 Th 232.038	59 Pa 231	60 U 238.029	61 Np 237	62 Pu 244	63 Am 243	64 Cm 247	65 Bk 247	66 Cf 251	67 Es 252	68 Fm 257	69 Md 258	70 No 259	71 Lr 262	72 Uub 263
73 Ta 180.948	74 W 183.84	75 Re 186.207	76 Os 190.23	77 Ir 192.222	78 Pt 195.084	79 Au 196.967	80 Hg 200.59	81 Tl 204.383	82 Pb 207.2	83 Bi 208.980	84 Po 209	85 At 210	86 Rn 222	87 Fr 223	88 Ra 226	89 Ac 227	90 Th 232.038
91 Pa 231	92 U 238.029	93 Np 237	94 Pu 244	95 Am 243	96 Cm 247	97 Bk 247	98 Cf 251	99 Es 252	100 Fm 257	101 Md 258	102 No 259	103 Lr 262	104 Uut 263	105 Uuq 264	106 Uub 265	107 Uuq 266	108 Uuo 267

1	ATOMIC NUMBER
H	SYMBOL
1.008	ATOMIC WEIGHT

- Elemental forms:**
  - Monoatomic:** Noble gases: helium, neon, argon, krypton, xenon, and radon; nonreactive
  - Diatomic:** Gases: fluorine, chlorine, hydrogen, oxygen, and nitrogen; Liquid: bromine; Solid: iodine
  - Molecular solids:** Sulfur
  - Covalent solids:** Carbon (diamond, graphite)
  - Metals:** Most common type of element; high coordination numbers; very dense
- Types of compounds:** Large variety due to ionic, covalent, and metallic bonding options
  - Ionic:** Electron exchange
  - Covalent:** Shared electrons
    - Polar covalent: Unequal sharing
  - Ion complex:** Central ion + Solvent
  - Transition metal compounds:**
    - Coordination compounds: Non-bonding ligands
    - Organometallics: Metal-C bond
  - Complex minerals:** Metal oxides and silicates
- Periodicity:** The chemical properties of an element depend on valence electrons
  - Family of elements:** Column in periodic table
  - The addition of d-orbitals produces an expanded octet

### Periodic Trends

- Chemistry determined by the behavior of valence electrons
- Atoms or ions get larger as you go down the column; electron shells are added
- Cation and anions (same charge) decrease in size as the atomic number ( $Z$ ) increases

- Ionization energies:** Noble gases are the most stable; alkali metals are the easiest to ionize; electron screening lessens  $Z$ -effect on orbital energy
- Atomic and ionic radii:** Not real properties; derive from structural data; divide bond into radii using a consistent scheme
  - Ionic radius:** Depends on charge, counter ion, and coordination (crystal type)
  - Atomic radius:** Based on diatomic gas data or metallic solid structural data

### Development of the Periodic Table

- Missing elements were predicted using periodicity
- Columns 1: Alkali metals
- Columns 2: Alkaline earth metals
- First transition series: Titanium-Copper
- Column 17: Halogens: fluorine, chlorine, bromine, iodine, and astatine
- Column 18: Rare gases: helium, neon, argon, krypton, xenon, and radon
- Artificial:  $Z = 90$  or above; radioactive
- Lanthanides: f-subshell elements; starts with  $Z = 58$



### Practical Application

- The colors of fireworks depend on the elements in the explosive

d-subshell  
5 orbitals



## Atomic Structure

- Atomic number ( $Z$ ):** The number of protons in the nucleus; for a neutral atom,  $Z$  = number of electrons; for an anion of charge  $-q$ ,  $Z = +q$  electrons; for a cation of charge  $+q$ ,  $Z = -q$  electrons
- Mass number ( $A$ ):**  $A = Z + N$ , where  $N$  = the number of neutrons in the nucleus
- Isotopes:** Atoms with the same  $Z$ , different  $A$ ; an element can have one or more isotopes; any sample contains a number of isotopes; in practical work, use the average mass number or atomic weight (weighted average of natural isotopes for a given element)

First Ionization Potential vs.  $Z$



Atomic Radius vs.  $Z$



## Atomic Quantum Numbers & Orbitals

- $n$ : principle**
  - $n$ : magnetic (orbital direction);  $n, l$  electron spin
- For each  $n$ , the possible  $l$  values are  $0, 1, \dots, n-1$
- For each  $l$ , the possible  $m_l$  values are  $-l, \dots, 0, \dots, +l$
- $m_s$  has two possible values:  $+\frac{1}{2}$  and  $-\frac{1}{2}$  (spin up, spin down)
- Each value of  $n$  denotes a "shell" in the atomic structure; each  $l$  denotes a subshell
  - s-orbital:  $l = 0$ , 1 type
  - d-orbital:  $l = 2$ , 5 types
  - p-orbital:  $l = 1$ , 3 types
  - f-orbital:  $l = 3$ , 7 types

### Many-Electron Atoms

Orbitals and quantum numbers derived for the hydrogen atom are used to describe all many-electron atoms and ions; a more rigorous treatment determines exact energy levels and orbital properties; the main result is that subshells have different energies; the filling of levels is guided by the Aufbau principle:

Aufbau Order for Filling Subshells



$n$	$l$	$m_l$	$m_s$	element
1	0	0	$+\frac{1}{2}$	H
1	0	0	$-\frac{1}{2}$	He
2	0	0	$+\frac{1}{2}$	Li
2	0	0	$-\frac{1}{2}$	Be
2	1	-1	$+\frac{1}{2}$	B
2	1	0	$+\frac{1}{2}$	C
2	1	1	$+\frac{1}{2}$	N
2	1	-1	$-\frac{1}{2}$	O
2	1	0	$-\frac{1}{2}$	F
2	1	1	$-\frac{1}{2}$	Ne

- Electron configuration:** Atomic orbital occupancy for an atom
- Pauli exclusion principle:** Each electron has a unique set of quantum numbers; an orbital may hold up to two electrons: one with spin up,  $m_s = +\frac{1}{2}$ , and one with spin down,  $m_s = -\frac{1}{2}$
- Hund's rule:** Electrons fill p, d, and f subshells in a manner to maximize the number of unpaired electrons (half-filled orbitals); spin is maximized
- Ionization potential (IP):** Energy required to remove electron from an atom or ion
  - First IP:** Removal of first valence electron
- Electronegativity:** Tendency of an atom to attract electrons in a chemical bond; 0-4 scale, 0 for rare gas

# Lavc Chemistry Study Guide

**Robert Rakes Shrock**



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