

# ACID-BASE RELATIVE STRENGTHS in Aqueous Solutions at 25° C

| Name of Acid               | Acid                                            |   | Base                                                                        |  | K <sub>a</sub>           |
|----------------------------|-------------------------------------------------|---|-----------------------------------------------------------------------------|--|--------------------------|
| Hydriodic                  | HI                                              | → | H <sup>+</sup> + I <sup>-</sup>                                             |  | >> 1                     |
| Hydrobromic                | HBr                                             | → | H <sup>+</sup> + Br <sup>-</sup>                                            |  | >> 1                     |
| Hydrochloric               | HCl                                             | → | H <sup>+</sup> + Cl <sup>-</sup>                                            |  | >> 1                     |
| Nitric                     | HNO <sub>3</sub>                                | → | H <sup>+</sup> + NO <sub>3</sub> <sup>-</sup>                               |  | > 1                      |
| Sulfuric                   | H <sub>2</sub> SO <sub>4</sub>                  | → | H <sup>+</sup> + HSO <sub>4</sub> <sup>-</sup>                              |  | > 1                      |
| Oxalic                     | H <sub>2</sub> C <sub>2</sub> O <sub>4</sub>    | ⇌ | H <sup>+</sup> + HC <sub>2</sub> O <sub>4</sub> <sup>-</sup>                |  | 5.90 × 10 <sup>-2</sup>  |
| Sulfurous                  | H <sub>2</sub> SO <sub>3</sub>                  | ⇌ | H <sup>+</sup> + HSO <sub>3</sub> <sup>-</sup>                              |  | 1.54 × 10 <sup>-2</sup>  |
| Hydrogen Sulfate Ion       | HSO <sub>4</sub> <sup>-</sup>                   | ⇌ | H <sup>+</sup> + SO <sub>4</sub> <sup>2-</sup>                              |  | 1.20 × 10 <sup>-2</sup>  |
| Phosphoric                 | H <sub>3</sub> PO <sub>4</sub>                  | ⇌ | H <sup>+</sup> + H <sub>2</sub> PO <sub>4</sub> <sup>-</sup>                |  | 7.52 × 10 <sup>-3</sup>  |
| Nitrous                    | HNO <sub>2</sub>                                | ⇌ | H <sup>+</sup> + NO <sub>2</sub> <sup>-</sup>                               |  | 4.6 × 10 <sup>-4</sup>   |
| Hydrofluoric               | HF                                              | ⇌ | H <sup>+</sup> + F <sup>-</sup>                                             |  | 3.53 × 10 <sup>-4</sup>  |
| Formic                     | HCO <sub>2</sub> H                              | ⇌ | H <sup>+</sup> + HCO <sub>2</sub> <sup>-</sup>                              |  | 1.77 × 10 <sup>-4</sup>  |
| Benzoic                    | C <sub>6</sub> H <sub>5</sub> CO <sub>2</sub> H | ⇌ | H <sup>+</sup> + C <sub>6</sub> H <sub>5</sub> CO <sub>2</sub> <sup>-</sup> |  | 6.46 × 10 <sup>-5</sup>  |
| Acetic                     | CH <sub>3</sub> CO <sub>2</sub> H               | ⇌ | H <sup>+</sup> + CH <sub>3</sub> CO <sub>2</sub> <sup>-</sup>               |  | 1.76 × 10 <sup>-5</sup>  |
| Carbonic                   | H <sub>2</sub> CO <sub>3</sub>                  | ⇌ | H <sup>+</sup> + HCO <sub>3</sub> <sup>-</sup>                              |  | 4.30 × 10 <sup>-7</sup>  |
| Hydrogen Sulfite Ion       | HSO <sub>3</sub> <sup>-</sup>                   | ⇌ | H <sup>+</sup> + SO <sub>3</sub> <sup>2-</sup>                              |  | 1.02 × 10 <sup>-7</sup>  |
| Dihydrogen Phosphate Ion   | H <sub>2</sub> PO <sub>4</sub> <sup>-</sup>     | ⇌ | H <sup>+</sup> + HPO <sub>4</sub> <sup>2-</sup>                             |  | 6.23 × 10 <sup>-8</sup>  |
| Boric                      | H <sub>3</sub> BO <sub>3</sub>                  | ⇌ | H <sup>+</sup> + H <sub>2</sub> BO <sub>3</sub> <sup>-</sup>                |  | 7.3 × 10 <sup>-10</sup>  |
| Ammonium Ion               | NH <sub>4</sub> <sup>+</sup>                    | ⇌ | H <sup>+</sup> + NH <sub>4</sub> <sup>+</sup>                               |  | 5.64 × 10 <sup>-10</sup> |
| Hydrogen Carbonate Ion     | HCO <sub>3</sub> <sup>-</sup>                   | ⇌ | H <sup>+</sup> + CO <sub>3</sub> <sup>2-</sup>                              |  | 5.61 × 10 <sup>-11</sup> |
| Monohydrogen Phosphate Ion | HPO <sub>4</sub> <sup>2-</sup>                  | ⇌ | H <sup>+</sup> + PO <sub>4</sub> <sup>3-</sup>                              |  | 2.2 × 10 <sup>-13</sup>  |
| Water                      | H <sub>2</sub> O                                | ⇌ | H <sup>+</sup> + OH <sup>-</sup>                                            |  | 1.0 × 10 <sup>-14</sup>  |

For space considerations, all reactions omit water as a reactant



# Lab 16 1 Creating Acid Base Strength Table

**Ensheng Dong**



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