

6. When one gram acid of KMnO_4 reacts with HCl , the volume of chlorine liberated at NTP will be
- (a) 11.2 litres (b) 22.4 litres
(c) 44.8 litres (d) 56.8 litres
7. When MnO_4^- ions reacts with Fe^{2+} ions in acid solution the Fe^{3+} is liberated. The equivalent weight of KMnO_4 in this reaction is
- (a) 500 (b) 1500
(c) 300 (d) 2000
8. 54 g. of hydrogen peroxide is present in 1000 ml. of solution. This solution is called
- (a) 10 vol solution (b) 20 vol solution
(c) 30 vol solution (d) 50 vol solution
9. The number of moles of KMnO_4 that will be needed to react completely with one mole of ferrous oxalate in acidic solution is
- (a) 10 (b) 20
(c) 40 (d) 2
10. To prepare a solution that is 0.05 M KCl starting with 100 mL of 0.40 M KCl
- (a) add 0.30 g KCl
(b) add 50 mL of water
(c) add 0.30 mol KCl
(d) evaporate 10 mL water
11. In acid solution solution, Br_2 disproportionates to Br^- and BrO_3^-
- $$3\text{Br}_2 + 6\text{H}_2\text{O} \rightarrow 2\text{Br}^- + \text{BrO}_3^- + 6\text{H}^+$$
- Atomic equivalent weight of Br_2 is (molecular weight = 160)
- (a) $\frac{16}{3}$ (b) $\frac{16}{2}$
(c) $\frac{160}{3}$ (d) $\frac{160}{2}$
12. (Fe^{2+}) in a solution prepared by adding 20.00 mL of 0.10 M FeSO_4 with 70 mL of 0.10 M Na_2SO_4 is
- (a) 0.020 M (b) 0.140 M
(c) 0.010 M (d) 0.200 M
13. When 50 mL of 0.100 M HCl is mixed with 100 mL of 0.10 M KOH , the resultant solution is the same as a solution of
- (a) 0.10 M HCl and 0.05 M HCl
(b) 0.05 M HCl
(c) 0.05 M HCl and 0.05 M KOH
(d) 0.05 M HCl and 0.05 M HCl
14. 50 mL of H_2O_2 solution contains strength = x% required 10 mL of KMnO_4 solution in acidic medium. Hence x is
- (a) 0.20 (b) 0.4
(c) 0.5 (d) 06.8
15. Molarity of 1.0 M H_2SO_4 (d = 1.8 mL/cm³) is
- (a) 36 mol kg⁻¹ (b) 300 mol kg⁻¹

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