

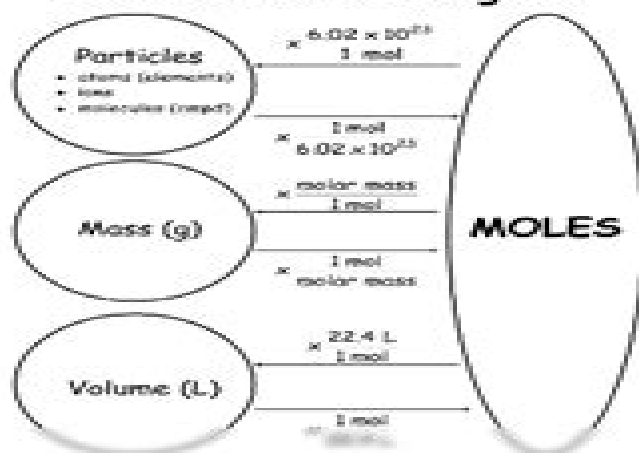
Honors Chemistry B- Final Exam Formula Sheet

Note: You will not use all of these formulas for this test

Temperature Formulas:

$$^{\circ}\text{C} = (^{\circ}\text{F} - 32) \cdot \frac{5}{9} \quad ^{\circ}\text{F} = \frac{9}{5}^{\circ}\text{C} + 32 \quad \text{K} = ^{\circ}\text{C} + 273$$

Mole Conversion Diagram



$$1 \text{ atm} = 760 \text{ mm Hg} = 101.3 \text{ kPa}$$

$$P_1 + P_2 = P_3 + P_4$$

$$\frac{Q}{T} = \frac{Q}{T}$$

$$\frac{Q}{T} = \frac{Q}{T}$$

$$\Delta G = \Delta H - T\Delta S$$

$$P \times T = n \times R \times T = PV = nRT$$

$$P_{\text{total}} = P_1 + P_2 + P_3 + \dots$$

$$\Delta G = \Delta H - T\Delta S$$

$$\text{Percent Yield} = \frac{\text{Actual Yield}}{\text{Theoretical Yield}} \times 100\%$$

$$M_1 \times V_1 = M_2 \times V_2$$

$$\text{Percent by Mass} = \frac{\text{Mass of Solute}}{\text{Mass of Solution}} \times 100$$

$$\text{Percent by Volume} = \frac{\text{Volume of Solute}}{\text{Volume of Solution}} \times 100$$

$$\text{Mole fractions: } X_A = \frac{n_A}{n_A + n_B} \quad X_B = \frac{n_B}{n_A + n_B}$$

$$1 \text{ Calorie} = 1 \text{ kilocalorie} = 1000 \text{ calories}$$

$$Q = mc\Delta T$$

$$q_{\text{sys}} = \Delta H = -q_{\text{surr}} = -m \times C \times \Delta T$$

$$\Delta H^{\circ} = \Delta H_f^{\circ}(\text{products}) - \Delta H_f^{\circ}(\text{reactants})$$

$$K_c = \frac{[\text{C}]^c \times [\text{D}]^d}{[\text{A}]^a \times [\text{B}]^b}$$

$$\Delta G = \Delta H - T\Delta S$$

$$\Delta G = \Delta H - T\Delta S$$

$$K_w = [\text{H}^+] \times [\text{OH}^-] = 1.0 \times 10^{-14}$$

$$\text{pH} = -\log[\text{H}^+]$$

$$\text{pOH} = -\log[\text{OH}^-]$$

$$K_c = \frac{[\text{H}^+][\text{A}^-]}{[\text{HA}]}$$

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JE Gale



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web the rate determining slower step gives a rate law showing second order dependence on the $[A]^2$ concentration and the sum of the two elementary equations gives the overall net reaction in general when the rate determining slower step is the first step in the reaction mechanism the rate law for the overall reaction is the same as the rate

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web the rate determining step limits the overall rate and therefore determines the rate law for the overall reaction introduction multistep reactions one of the most important applications of kinetics is to the study of reaction mechanisms or the sequences of steps by which chemical reactions occur

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web the initial rate law is then measured for each of the reactions consider the reaction between nitrogen monoxide gas and hydrogen gas to form nitrogen gas and water vapor the following data were collected for this reaction at 1280 text o text c see table below notice that the starting concentrations of ce no and ce h

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web determination of a rate law and activation energy in this experiment glass beakers pipettes and a hotplate were utilized to set up the procedures a solution of deionized water a buffer potassium iodide ki sodium thiosulfate $\text{na}_2\text{S}_2\text{O}_3$ starch and hydrogen peroxide H_2O_2 were used in order to calculate the rate of reaction

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web oct 27 2022 the rate law $\text{rate } k [\text{H}_2\text{O}_2]$ describes a reaction that is first order in H_2O_2 first order in OH^- and second order overall example 12 3 1 writing rate laws from reaction orders an experiment shows that the reaction of

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web rate laws sometimes called differential rate laws or rate equations are mathematical expressions that describe the relationship between the rate of a chemical reaction and the concentration of its reactants as an example consider the

reaction described by the chemical equation $a + b \rightarrow \text{products}$

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