

**Nuclear Reactions**

Nuclear reactions are characterized by the following:

- They involve changes to the nucleus
  - Number of protons or neutrons in the nucleus change
- Energy changes are enormous
- Not concerned with valence electrons.
- Changes in temperature and pressure have no effect on the rate of the reaction
- Small particles are important.

**Important subatomic particles**

proton:  ${}^1_1\text{p}$      ${}^1_1\text{H}$

neutron:  ${}^1_0\text{n}$

Alpha Particle:  ${}^4_2\alpha$      ${}^4_2\text{He}$     2 protons  
2 neutrons

beta particle:  ${}^0_{-1}\beta$      ${}^0_{-1}\text{e}$     - high energy electrons emitted from the nucleus

positron:  ${}^0_{+1}\beta$      ${}^0_{+1}\text{e}$     - positive electron: what gives a proton its positive charge

Gamma Ray:  ${}^0_0\gamma$     - high energy photon

**Definitions:**

**Isotopes** - same number of protons but different mass number (mass number = # protons + # neutrons)

${}^{12}_6\text{C}$      ${}^{13}_6\text{C}$      ${}^{14}_6\text{C}$     are each an isotope of C (How many neutrons does each have?)

**Nuclide** (pronounced NOO clyde) - a nucleus with a specific number of protons and neutrons

**Radioactive** - Changes in matter resulting from SPONTANEOUS changes in the composition of the nucleus.

- Such changes are often accompanied by the creation of the subatomic particles above.
- Changes to the nucleus often result in the conversion of one element into another.

**Radionuclide** - nuclides that are radioactive.



# Guide Nuclear Chemistry Answers

**L Reisser**



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