

Your Driving Test Checklist

Pass your theory test



Check your vision

- read new-style plate from 20m away
- read old-style plate from 20.5m away

AF01 REG



Know the show me questions

- show how to wash/clean the rear windscreen
- show how to wash/clean the front windscreen
- show how to switch on your dipped headlights
- show how to set the rear demister
- show how to operate the horn
- show how to demist the front windscreen
- show how to open & close the side window



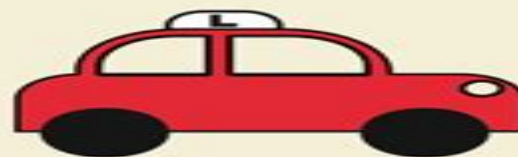
Know the tell me questions

- explain how to check the brakes are working
- explain where your tyre pressure info is & how to check it
- explain how to ensure your head restraint is correctly adjusted
- explain how to check tyres have sufficient thread/are in safe condition
- explain how to check headlights & tail lights work
- explain how you'd know if there was a problem with your ABS
- explain how you'd make sure the direction indicators work
- explain how to check that the brake lights work
- explain how to check the power-assisted steering is working
- explain how to switch on the rear fog light(s) & when to use them
- explain how to switch your headlights from dipped to main beam & how you'd know the main beam is on
- OPEN THE BONNET & explain how you'd check the engine oil
- OPEN THE BONNET & explain how you'd check the coolant
- OPEN THE BONNET & explain how you'd check you have a safe level of hydraulic brake fluid



Know the manoeuvres

- Pulling up to the right
- Bay parking
- Parallel parking
- Emergency stop



Practise general driving

- Be able to pull over & away in these situations:
 - Stopping at the side of the road
 - Pulling over & pulling out from behind a parked vehicle
 - Pulling over on a hill and performing a hill start
- Practise most of the routes in your test area



Practise independent driving

- Feel comfortable following a satnav
- Feel comfortable following traffic signs



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web with a strong acid and strong base titration strong acids and strong bases refer to species that completely dissociate to form ions in solution by contrast weak acids and bases ionize only partially and the ionization reaction is reversible

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web aug 30 2022 acid base titrations are usually used to find the amount of a known acidic or basic substance through acid base reactions the analyte titrand is the solution with an unknown molarity the reagent titrant is the solution with a known molarity that will react with the analyte

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web conclusion the purpose of the experiment was to identify the unknown solid acid 17 by calculating its molar mass this was done by titration of the acid with sodium hydroxide in order to find the volume of sodium hydroxide required to reach the end point and that value was used in unit analysis to find the molar mass and identify the unknown solid acid

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web by utilizing the half equivalence point on the titration curve the pka of the acid analyte was determined to be 3.655 furthermore using the stoichiometric ratio between the naoh and the acid analyte 1:1 we could experimentally determine the number of moles of the acid in the solution to be 0.00731 moles making the experimental molar mass

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web chem 1210 spring 2019 experiment 10 11 part 1 acid base titration abstract the purpose of this experiment is to observe the titration of hydrochloric acid a strong acid with sodium hydroxide a strong base and acetic acid a

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web goals understand that parts of an acid base titration be able to determine the K_a or K_b from pH data associated with the titration of a weak acid or base be able to determine the molar mass of a solid monoprotic acid from titration data be able to calculate K_{a1} and K_{a2} for a polyprotic acid by the end of this lab students should be able to

21 17 titration experiment chemistry libretexts - Oct 06 2022

web titration experiment in the neutralization of hydrochloric acid by sodium hydroxide the mole ratio of acid to base is 1 1
 $HCl(aq) + NaOH(aq) \rightarrow NaCl(aq) + H_2O(l)$
 one mole of HCl would be fully neutralized by one mole of NaOH

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web conclusion the identification of an unknown solid acid the purpose of this experiment was to determine the standardization of sodium hydroxide NaOH and oxalic acid $H_2C_2O_4 \cdot 2H_2O$ and the molar mass of an unknown solid acid using the molar mass of the standardized sodium hydroxide solution

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web acid base reactions can be used to determine concentrations of an acid or a base using stoichiometric calculations and titrimetric analysis the volume of a solution required to react

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web jul 19 2023 introduction titration is an analytical quantitative technique used to determine the concentration of a solute a pH titration is used to determine the concentration of an acid or a base titrations play an important role in determining amount and purity in many manufacturing processes

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web specifically an acid base titration can be used to figure out the concentration of an acid or base whether an unknown acid or base is strong or weak pK_a of an unknown acid or pK_b of the unknown base key terms titration a process where a solution of known strength is added to a certain volume of a treated sample containing an

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web conclusion this lab evaluated an acid base reaction with the volumetric analysis of a standard solution being NaOH reaction with a 2nd substance $H_2C_2O_4(aq) + 2NaOH(aq) \rightarrow Na_2C_2O_4(aq) + 2H_2O(l)$ to differentiate an acid from a base in a solution a pH scale is used the pH scale ranges from 0 to 14 a pH of 7 is neutral in an acid base reaction

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web conclusion in this lab we used titration to explore the concepts of stoichiometry and equivalence points we found the concentration of an unknown substance by mixing 2m hcl with the naoh of unknown concentration in order to experimentally ascertain the concentration of the naoh

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web this point of titration is referred to as the equivalence point considering the 1 1 stoichiometry of this acid base reaction
 $\text{NaOH (aq)} + \text{C}_6\text{H}_4\text{COOH (aq)} \rightarrow \text{C}_6\text{H}_4\text{COONa (aq)} + \text{H}_2\text{O (l)}$
the point of equivalence is the point of titration when the number of moles of NaOH added is equal to the number of moles of KHP in the