

### Types of Chemical Reactions Worksheet #1

#### **11-Chemistry**

For each of the following equations, identify what kind of reaction it represents: double replacement, single replacement, decomposition, or synthesis (composition).

- 1)  $2\text{Mg} + \text{O}_2 \rightarrow 2\text{MgO}$  **Synthesis**
- 2)  $\text{CaI}_2 + \text{Cl}_2 \rightarrow \text{CaCl}_2 + \text{I}_2$  **Single Replacement**
- 3)  $\text{HClOH} + \text{AlCl}_3 \rightarrow \text{AlOHCl}_2 + 2\text{HCl}$  **Double Replacement**
- 4)  $\text{C} + \text{O}_2 \rightarrow \text{CO}_2$  **Synthesis**
- \*5)  $\text{CaOHCl}_2 + 2\text{HCl} \rightarrow \text{CaCl}_2 + 2\text{H}_2\text{O}$  **Double Replacement**
- 6)  $2\text{KClO}_3 \rightarrow 2\text{KCl} + 3\text{O}_2$  **Decomposition**
- 7)  $3\text{CuSO}_4 + 2\text{Al} \rightarrow \text{Al}_2(\text{SO}_4)_3 + 3\text{Cu}$  **Single Replacement**
- 8)  $\text{Na}_2\text{S} + 2\text{AgNO}_3 \rightarrow 2\text{NaNO}_3 + \text{Ag}_2\text{S}$  **Double Replacement**
- 9)  $3\text{H}_2 + \text{N}_2 \rightarrow 2\text{NH}_3$  **Synthesis**

\*10)  $2\text{Na} + 2\text{H}_2\text{O} \rightarrow 2\text{NaOH} + \text{H}_2$  **Single Replacement**

\*Remember: " $\text{H}_2\text{O}$ " can also be thought of as " $\text{HOH}$ "

Complete each of the following equations as needed to make it the type of reaction indicated. Be sure to write each formula correctly.

- 11) Double replacement:  $\text{Na}_2\text{CrO}_4 + \text{PbCl}_2 \rightarrow 2\text{NaCl} + \text{PbCrO}_4$
- 12) Single replacement:  $\text{Cl}_2 + 2\text{NaBr} \rightarrow 2\text{NaCl} + \text{Br}_2$
- 13) Decomposition:  $\text{Mg}(\text{ClO}_3)_2 \rightarrow \text{MgCl}_2 + 3\text{O}_2$
- 14) Synthesis:  $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$
- 15) Double replacement:  $3\text{CaOHCl}_2 + 2\text{FeCl}_3 \rightarrow 2\text{FeOHCl}_2 + 3\text{CaCl}_2$
- 16) Single replacement:  $\text{Fe} + \text{Cu}(\text{NO}_3)_2 \rightarrow \text{Fe}(\text{NO}_3)_2 + \text{Cu}$  [Assume  $\text{Fe}^{2+}$ ]
- 17) Decomposition:  $2\text{Hg}_2\text{O} \rightarrow 4\text{Hg} + \text{O}_2$
- 18) Synthesis:  $\text{S} + \text{O}_2 \rightarrow \text{SO}_2$
- 19) Double replacement:  $\text{AgNO}_3 + \text{KI} \rightarrow \text{AgI} + \text{KNO}_3$
- 20) Single replacement:  $\text{Cu} + 2\text{AgNO}_3 \rightarrow 2\text{Ag} + \text{Cu}(\text{NO}_3)_2$  [Copper (II) is used here]

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