

Terms to Know:

- Electrochemistry**—the study of the interchange of chemical and electrical energy
OIL RIG – oxidation is loss, reduction is gain (of electrons)
Oxidation – the loss of electrons, increase in charge
Reduction – the gain of electrons, reduction of charge
Oxidation number – the assigned charge on an atom
Oxidizing agent (OA) – the species that is reduced and thus *causes* oxidation
Reducing agent (RA) – the species that is oxidized and thus *causes* reduction

ELECTROCHEMISTRY INVOLVES TWO MAIN TYPES OF PROCESSES:

- Galvanic (voltaic) cells** – which are spontaneous chemical reactions (battery)
- Electrolytic cells** – which are non-spontaneous and require external e^- source (DC power source)
- BOTH** of these fit into the category entitled **Electrochemical cells**

GALVANIC CELLS

Parts of the voltaic or galvanic cell:

- Anode--the electrode where oxidation occurs. After a period of time, the anode may appear to become smaller as it falls into solution.
- Cathode-- the anode where reduction occurs. After a period of time it may appear larger, due to ions from solution plating onto it.
- inert electrodes—used when a gas is involved OR ion to ion involved such as Fe^{3+} being reduced to Fe^{2+} rather than Fe^0 . Made of Pt or graphite.
- Salt bridge -- a device used to maintain electrical neutrality in a galvanic cell. This may be filled with agar which contains a neutral salt or it may be replaced with a porous cup.
- Electron flow -- always from anode to cathode. (through the wire)
- Standard cell notation (line notation) - anode/solution// cathode solution/ cathode Ex. $Zn/Zn^{2+} (1.0 M) // Cu^{2+} (1.0 M) / Cu$
- Voltmeter - measures the cell potential (*emf*) . Usually is measured in volts.

Balance this re-dox reaction:



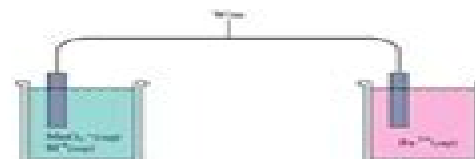
RED:

OA:

OX:

RA

Overall rxn:



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