

20 Balanced Chemical Equations

1. $2\text{Na (s)} + 2\text{H}_2\text{O (l)} \rightarrow 2\text{NaOH (aq)} + \text{H}_2 \text{ (g)}$
2. $\text{Zn(s)} + \text{H}_2\text{SO}_4 \text{ (aq)} \rightarrow \text{ZnSO}_4 \text{ (aq)} + \text{H}_2 \text{ (g)}$
3. $\text{CaC}_2 \text{ (s)} + 2\text{H}_2\text{O (l)} \rightarrow \text{C}_2\text{H}_2 \text{ (g)} + \text{Ca(OH)}_2 \text{ (aq)}$
4. $2\text{NO (g)} + \text{O}_2 \text{ (g)} \rightarrow 2\text{NO}_2 \text{ (g)}$
5. $\text{Mg (s)} + 2\text{H}_2\text{O (l)} \rightarrow \text{Mg(OH)}_2 \text{ (aq)} + \text{H}_2 \text{ (g)}$
6. $3\text{Fe (s)} + 4 \text{H}_2\text{O (g)} \rightarrow \text{Fe}_3\text{O}_4 \text{ (g)} + 4\text{H}_2 \text{ (g)}$
7. $\text{CH}_4 \text{ (g)} + \text{H}_2\text{O (g)} \rightarrow \text{CO (g)} + 3\text{H}_2 \text{ (g)}$
8. $\text{CuO (s)} + \text{H}_2 \text{ (g)} \rightarrow \text{Cu (s)} + \text{H}_2\text{O (l)}$
9. $3\text{H}_2 \text{ (g)} + \text{N}_2 \text{ (g)} \rightarrow 2\text{NH}_3 \text{ (g)}$
10. $\text{Pb (s)} + 4\text{Na (s)} + 4\text{C}_2\text{H}_5\text{Cl (g)} \rightarrow 4\text{NaCl (s)} + \text{Pb(C}_2\text{H}_5)_4 \text{ (l)}$
11. $\text{TiCl}_4 \text{ (l)} + 4\text{Na (s)} \rightarrow 4\text{NaCl (s)} + \text{Ti (s)}$
12. $2\text{NaOH (aq)} + \text{CO}_2 \text{ (g)} \rightarrow \text{Na}_2\text{CO}_3 \text{ (aq)} + \text{H}_2\text{O (l)}$
13. $\text{Cl}_2 \text{ (g)} + 6\text{NaOH (aq)} \rightarrow \text{NaClO}_3 \text{ (aq)} + 5\text{NaCl (aq)} + 3\text{H}_2\text{O (l)}$ (for concentrated aqueous NaOH)
14. $\text{Na}_2\text{SiO}_3 \text{ (s)} + 5\text{SiO}_2 \text{ (s)} \rightarrow \text{Na}_6\text{Si}_6\text{O}_{18} \text{ (s)} + 5\text{O}_2 \text{ (g)}$
15. $\text{Na}_2\text{SiO}_3 \text{ (s)} + 4\text{C (s)} \rightarrow \text{Na}_2\text{S (s)} + 4\text{CO (g)}$
16. $\text{NaCl (aq)} + \text{NH}_4\text{OH (aq)} + \text{CO}_2 \text{ (g)} \rightarrow \text{NaHCO}_3 \text{ (s)} + \text{NH}_4\text{Cl (aq)}$
17. $2\text{NaHCO}_3 \text{ (s)} \rightarrow \text{Na}_2\text{CO}_3 \text{ (s)} + \text{H}_2\text{O} + \text{CO}_2 \text{ (g)}$
18. $\text{Na}_2\text{CO}_3 \text{ (s)} + 10\text{H}_2\text{O} \rightarrow \text{Na}_2\text{CO}_3 \cdot 10\text{H}_2\text{O (s)}$
19. $\text{Na}_2\text{CO}_3 \text{ (aq)} + \text{Ca(OH)}_2 \text{ (aq)} \rightarrow 2\text{NaOH (aq)} + \text{CaCO}_3 \text{ (s)}$
20. $\text{Fe}_3\text{O}_4 \text{ (s)} + 3\text{C (s)} \rightarrow 2\text{Fe (s)} + 3\text{CO (g)}$
21. $2\text{Cu(NO}_3)_2 \text{ (s)} \rightarrow 2\text{CuO (s)} + 4\text{NO}_2 \text{ (g)} + \text{O}_2 \text{ (g)}$
22. $2\text{Al (s)} + 6\text{HCl (aq)} \rightarrow 2\text{AlCl}_3 \text{ (aq)} + 3\text{H}_2 \text{ (g)}$
23. $2\text{Al (s)} + 2\text{NaOH (aq)} + 6\text{H}_2\text{O (l)} \rightarrow 2\text{NaAl(OH)}_4 \text{ (aq)} + 3\text{H}_2 \text{ (g)}$

Lab 20 Balanced Chemical Equations

J Rink



Lab 20 Balanced Chemical Equations:

Lab 20 Balanced Chemical Equations Book Review: Unveiling the Power of Words

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