

# INORGANIC CHEMISTRY



## CHEMICAL BONDING

### IONIC BONDS

- Atoms with low ionization energy form cations (1-3), and atoms with high electron affinity form anions (1-3) attracted to form ionic bonds.
- Approximate charges on cations and anions create an ionic bond.

### THE LEWIS OCTET RULE: COVALENT BONDS

- Atoms other than hydrogen (stable with 2 valence electrons) will form bonds until they are surrounded by 8 valence electrons, and have achieved stable noble gas configuration.
- In a single covalent bond, a pair of electrons is shared between 2 atoms. In multiple covalent bonds, 2 or 3 pairs are shared.



### BOND DISSOCIATION ENERGIES: CALCULATING $\Delta H$

- $\Delta H = \sum \text{Bond energies of reactants} - \sum \text{Bond energies of products}$
- Example:  $\text{CH}_4 + 2\text{O}_2 \rightarrow \text{CO}_2 + 2\text{H}_2\text{O}$
- $\Delta H = 4(413.4) + 2(495.8) - 2(463.1) - 2(517.9) = -802.6 \text{ kJ/mol}$
- Reaction is exothermic = energy (heat) flows out of system.
- Endothermic = energy (heat) flows into system.

### PERCENT IONIC CHARACTER OF A CHEMICAL BOND

- Bonds with 50% or greater ionic character are ionic.
- Bonds with less than 50% ionic character are covalent.
- Transition from ionic to covalent bonding results in a reduction in electrical conductivity, melting point, and boiling point.



## INTRODUCTION

- The study of compounds (group of carbon, except for  $\text{CO}$ ,  $\text{CO}_2$ ,  $\text{CS}_2$ , and compounds containing  $\text{CN}^-$ ,  $\text{CO}_3^{2-}$ , and  $\text{HCO}_3^-$ ).

### INORGANIC COMPOUNDS

Name	Formula	Name	Formula
Alumina	$\text{Al}_2\text{O}_3$	Glauber's Salt	$\text{Na}_2\text{SO}_4$
Ammonia	$\text{NH}_3$	Cyanate	$\text{CaNCN} \cdot 2\text{H}_2\text{O}$
Anhydrous	$\text{SO}_3$	Heavy Water	$\text{D}_2\text{O}$
Baking soda	$\text{NaHCO}_3$	Litharge	$\text{PbO}$
Bauxite	$\text{Al}_2\text{O}_3 \cdot \text{H}_2\text{O} \cdot \text{H}_2\text{O}$	Salt or Magnesia	$\text{MgO}$
Calcite	$\text{CaCO}_3$	Quartzite	$\text{CaSi}$
Carbon	$\text{H}_2\text{CO}_3$	Silica	$\text{SiO}_2$
Cyanide	$\text{Na}_3\text{CN}$	Sulphur Lime	$\text{CaSO}_4$
Dry Ice	$\text{CO}_2$ (solid)	Table Salt	$\text{NaCl}$
Gypsum	$\text{CaSO}_4 \cdot 2\text{H}_2\text{O}$	Washing Soda	$\text{Na}_2\text{CO}_3 \cdot 10\text{H}_2\text{O}$

### COMMON INORGANIC IONS

Ammonium ( $\text{NH}_4^+$ )	Mercuric ( $\text{Hg}_2^{2+}$ )	Nitride ( $\text{N}^{3-}$ )
Ammonium ( $\text{NH}_4^+$ )	Mercurous ( $\text{Hg}_2^{2+}$ )	Hydrogen Carbonate or Bicarbonate ( $\text{HCO}_3^-$ )
Barium ( $\text{Ba}^{2+}$ )	Potassium ( $\text{K}^+$ )	Hydrogen Phosphate ( $\text{HPO}_4^{2-}$ )
Calcium ( $\text{Ca}^{2+}$ )	Silver ( $\text{Ag}^+$ )	Hydrogen Sulfate, Bisulfate ( $\text{HSO}_4^-$ )
Cesium ( $\text{Cs}^+$ )	Sodium ( $\text{Na}^+$ )	Nitrate ( $\text{NO}_3^-$ )
Chromium ( $\text{Cr}^{3+}$ )	Strontium ( $\text{Sr}^{2+}$ )	Fluoride ( $\text{F}^-$ )
Cobalt ( $\text{Co}^{2+}$ )	Tin ( $\text{Sn}^{2+}$ )	Selenate ( $\text{SeO}_4^{2-}$ )
Copper ( $\text{Cu}^{2+}$ )	Zinc ( $\text{Zn}^{2+}$ )	Selenite ( $\text{SeO}_3^{2-}$ )
Copper ( $\text{Cu}^{2+}$ )	Bromide ( $\text{Br}^-$ )	Hydroxide ( $\text{OH}^-$ )
Copper ( $\text{Cu}^{2+}$ )	Carbonate ( $\text{CO}_3^{2-}$ )	Chloride ( $\text{Cl}^-$ )
Hydrogen ( $\text{H}^+$ )	Chloride ( $\text{ClO}_2^-$ )	Permanganate ( $\text{MnO}_4^-$ )
Iron ( $\text{Fe}^{2+}$ )	Chloride ( $\text{Cl}^-$ )	Formate ( $\text{HCO}_2^-$ )
Iron ( $\text{Fe}^{3+}$ )	Chlorate ( $\text{ClO}_3^-$ )	Phosphate ( $\text{PO}_4^{3-}$ )
Lead ( $\text{Pb}^{2+}$ )	Cyanide ( $\text{CN}^-$ )	Sulfate ( $\text{SO}_4^{2-}$ )
Lithium ( $\text{Li}^+$ )	Dichromate ( $\text{Cr}_2\text{O}_7^{2-}$ )	Sulfite ( $\text{SO}_3^{2-}$ )
Magnesium ( $\text{Mg}^{2+}$ )	Ethylhydrogen Phosphate ( $\text{C}_2\text{H}_5\text{HPO}_4^-$ )	Sulfate ( $\text{SO}_4^{2-}$ )
Magnesium ( $\text{Mg}^{2+}$ )	Bromide ( $\text{Br}^-$ )	Thiosulfate ( $\text{S}_2\text{O}_3^{2-}$ )

### PAULI'S RULE

- Indicates that one orbital has the same energy, electron occupancy, and spin with available spins.
- Using 2p orbitals to demonstrate.

### ORBITAL TYPES & SHAPES



### PAULI'S EXCLUSION PRINCIPLE

- No 2 electrons in an atom can have the same 4 quantum numbers. No more than 2 electrons can occupy any 1 orbital.

# Inorganic Elements Student Guide

**Robert R. Redfield**



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