

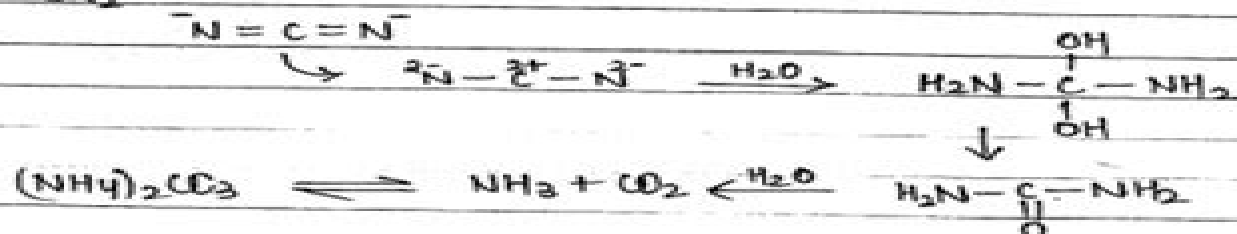
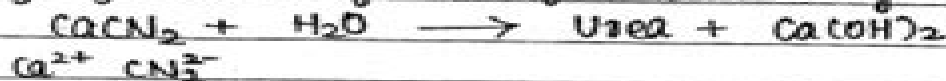
• N_2 is absorbed by CaC_2 at $1000^\circ C$
 $CaC_2 + N_2 \xrightarrow{1000^\circ C} CaCN_2 + C$

Ammonia (NH_3):

Preparation:

- ① Any NH_4^+ salt + Base $\rightarrow NH_3 \uparrow$
 $NH_4Cl + NaOH \rightarrow NH_3 \uparrow + NaCl$
- ② Any metal nitride + H_2O
 $Mg_3N_2 + H_2O \rightarrow Mg(OH)_2 + NH_3 \uparrow$
- ③ Heating of NH_4^+ salt having non/weak oxidising anion
 $(NH_4)_2CO_3 \xrightarrow{\Delta} NH_3 + H_2CO_3$
 $NH_4Cl \xrightarrow{\Delta} NH_3 + HCl$
 $NH_4CN \xrightarrow{\Delta} NH_3 + HCN$
- ④ $NO_2^- + \text{Fe} + NaOH \rightarrow NH_3 + NaNO_2$
 $\quad \quad \quad \uparrow$
 $\quad \quad \quad H_2$
- ⑤ $NO_3^- + Zn + NaOH \rightarrow NH_3 + Na_2ZnO_2$
- ⑥ Haber's process:
 $N_2 + 3H_2 \xrightarrow[\text{high pressure}]{Fe, Mo} 2NH_3 \quad \Delta H = -84$

Uses: ① Used in identification of transition metal cation.
 ② In preparation of nitrogen containing fertilizers.



Inorganic Chemistry Notes

F. H. Jeffery



Inorganic Chemistry Notes:

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