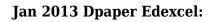
Acceptable Answers	Reject	Mark
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[HCOOH(aq)] = $\frac{1.50 \times 10^{-4}}{0.0250}$ OR = $1.50 \times 10^{-4} \times \frac{1000}{25.0}$		
$(= 6.00 \times 10^{-3} \text{ mol dm}^{-3})$ ALTERNATI VE APPROACH: Use of an expression such as 0.00750 x 20.0 = 25 x y (1) $y = 0.00750 \times 20.0$ 25		
	Reject	Mark
(Weak) dissociates / ionizes to a small extent OR dissociate / ionizes partially OR dissociates / ionizes incompletely OR does not fully dissociate / ionize OR forms an equilibrium when reacted with water (1) (Acid) proton donor ALLOW 'proton donator' OR produces / releases H* ions OR produces / releases H ₃ O* ions [1] Ignore reference to typical acid	'not easily dissociated'	2
	First mark Moles NaOH = 0.00750 x 20.0 1000 = 1.50 x 10 ⁻⁴ (mol) (1) (Since HCOOH: NaOH ratio is 1:1) Second mark [HCOOH(aq)] = 1.50 x 10 ⁻⁴ 0.0250 OR = 1.50 x 10 ⁻⁴ x 1000 25.0 (1) (= 6.00 x 10 ⁻³ mol dm ⁻³) ALTERNATI VE APPROACH: Use of an expression such as 0.00750 x 20.0 = 25 x y (1) y = 0.00750 x 20.0 25 (1) Acceptable Answers (Weak) dissociates / ionizes to a small extent OR dissociate / ionizes partially OR dissociates / ionizes incompletely OR does not fully dissociate / ionize OR forms an equilibrium when reacted with water (1) (Acid) proton donor ALLOW 'proton donator' OR produces / releases H* ions	First mark Moles NaOH = 0.00750 x 20.0 1000 = 1.50 x 10 ⁻⁴ (mol) (Since HCOOH: NaOH ratio is 1:1) Second mark [HCOOH(aq)] = 1.50 x 10 ⁻⁴ 0.0250 OR = 1.50 x 10 ⁻⁶ x 1000 25.0 (1) (= 6.00 x 10 ⁻³ mol dm ⁻³) ALTERNATI VE APPROACH: Use of an expression such as 0.00750 x 20.0 = 25 x y (1) y = 0.00750 x 20.0 25 (1) Acceptable Answers (Weak) dissociates / ionizes to a small extent OR dissociate / ionizes partially OR dissociates / ionizes incompletely OR does not fully dissociate / ionize OR forms an equilibrium when reacted with water (1) (Acid) proton donor ALLOW 'proton donator' OR produces / releases H ₂ O' ions Ignore reference to typical acid

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