

Quiz 4 – Potential Energy Diagrams

1. The addition of a catalyst to a reaction provides an alternative mechanism with

- A. Lower activation energy and lower reaction rate
- B. Lower activation energy and higher reaction rate**
- C. Higher activation energy and lower reaction rate
- D. Higher activation energy and higher reaction rate

2. Consider the following Reaction



The correct equation including the heat term is

- A. $\text{N}_{2(g)} + \text{O}_{2(g)} + 90 \text{ kJ} \longrightarrow 2\text{NO}_{(g)}$
- B. $\text{N}_{2(g)} + \text{O}_{2(g)} + 180 \text{ kJ} \longrightarrow 2\text{NO}_{(g)}$**
- C. $\text{N}_{2(g)} + \text{O}_{2(g)} \longrightarrow 2\text{NO}_{(g)} + 90 \text{ kJ}$
- D. $\text{N}_{2(g)} + \text{O}_{2(g)} \longrightarrow 2\text{NO}_{(g)} + 180 \text{ kJ}$

3. A forward reaction has an activation energy of 50 kJ and a ΔH of -100 kJ .
The PE diagram, which describes this reaction, is

A.

4. Consider the following potential energy diagram

The Activation energy for the forward reaction is

- A. 25 kJ
- B. 50 kJ
- C. 75 kJ**
- D. 125 kJ

5. Consider the following reaction



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Marco Cascella

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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Services, 2009-02-09 The book itself contains chapter length subject reviews on every subject tested on the AP Chemistry exam as well as both sample multiple choice and free response questions at each chapter's end Two full length practice tests with detailed answer explanations are included in the book **Problems and Problem Solving in Chemistry Education**

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