

SOLVED QUESTIONS

Multiple Choice Questions

1. K_1 and K_2 are the velocity constants of forward and backward reactions. The equilibrium constant k of the reaction is
- (A) $K_1 \times K_2$ (B) $K_1 - K_2$
 (C) $\frac{K_1}{K_2}$ (D) $\frac{K_1 + K_2}{K_1 - K_2}$
- Sol.** (C).
2. Which of the following reactions will be favoured at low pressure?
- (A) $N_2 + O_2 \rightleftharpoons 2NO$ (B) $H_2 + I_2 \rightleftharpoons 2HI$
 (C) $PCl_5 \rightleftharpoons 2PCl_3 + Cl_2$ (D) $N_2 + 3H_2 \rightleftharpoons 2NH_3$
- Sol.** (C).
3. For the reaction, $A + B \longrightarrow C + D$, if active mass of A is kept constant and active mass of B is tripled, the rate of reaction will become
- (A) three times (B) six times
 (C) eight times (D) nine times
- Sol.** (D).
4. The rate of $\frac{K_p}{K_c}$ for the reaction:
- $$CO(g) + \frac{1}{2}O_2(g) \rightleftharpoons CO_2(g)$$
- is
- (A) 1 (B) RT
 (C) $(RT)^{1/2}$ (D) $(RT)^{-1/2}$
- Sol.** (D).
5. Consider the reaction,
- $$CaCO_3(s) \rightleftharpoons CaO(s) + CO_2(g)$$
- in a closed container at equilibrium. What would be the effect of addition of $CaCO_3$ on the equilibrium concentration of CO_2 ?
- (A) Increases (B) Unpredictable
 (C) Decreases (D) Remains unaffected
- Sol.** (D).
6. For the reaction, $PCl_5(g) \rightleftharpoons PCl_3(g) + Cl_2(g)$, the forward reaction at constant temperature is favoured by
- (A) introducing an inert gas at constant volume
 (B) introducing $PCl_3(g)$ at constant volume
 (C) introducing $PCl_5(g)$ at constant volume
 (D) introducing $Cl_2(g)$ at constant volume
- Sol.** (C).
7. A liquid is in equilibrium with its vapour at its boiling point. On an average, the molecules in the two phases have equal
- (A) intermolecular forces
 (B) potential energy
 (C) kinetic energy
 (D) both kinetic energy and potential energy
- Sol.** (C).
8. The unit of equilibrium constant for the reversible reaction

Multiple Choice Questions In Chemical Equilibrium

Dr. Manish Rannjan (IAS)



Multiple Choice Questions In Chemical Equilibrium:

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