

Questions and Problems

Q1 According to your graph, what is the relationship between the pressure and the volume of a gas?

* As the volume of the gas increases the pressure of the gas decreases. They have inverse proportional relationship.

Q2 On your graph, what is the volume of the gas at a pressure of 760. mmHg?

* After plotting the data to excel, the graph produced a linear equation $y = -28.873x + 1539.3$ where y is the pressure and x is the volume of gas.

We need to find x at 760 mmHg.

Hence,

$$760 \text{ mmHg} = -28.873x + 1539.8 \Rightarrow x = 26.99 \text{ mL}$$

Therefore, at 760 mmHg the volume of the gas is 26.99 mL

Q3 On your graph, what is the pressure of the gas when the volume is 30.0 mL?

* $x = 30.0 \text{ mL}$, $y = ?$

$$y = -28.873 (30.0) + 1539.3$$

$$y = 673.1 \text{ mmHg}$$

Therefore, when the volume of the gas is 30.0 mL the pressure is 673.1 mmHg

Q4 A sample of helium has a volume of 325 mL and a pressure of 655 mmHg. What will be the pressure, in mmHg, if the sample of helium is compressed to 125 mL. (T, n constant)?

$$P_1 V_1 = P_2 V_2$$

$$P_2 = (P_1 V_1) \div V_2 = (655 \text{ mmHg} \times 325 \text{ mL}) \div 125 \text{ mL} = 1,703 \text{ mmHg}$$

Holt Chemistry Section Review Answers Gases

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In a fast-paced digital era where connections and knowledge intertwine, the enigmatic realm of language reveals its inherent magic. Its capacity to stir emotions, ignite contemplation, and catalyze profound transformations is nothing short of extraordinary. Within the captivating pages of **Holt Chemistry Section Review Answers Gases** a literary masterpiece penned by way of a renowned author, readers set about a transformative journey, unlocking the secrets and untapped potential embedded within each word. In this evaluation, we shall explore the book's core themes, assess its distinct writing style, and delve into its lasting effect on the hearts and minds of people who partake in its reading experience.

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