

## **Grade 11      November Exam Scope**

### **Mathematics:**

#### **Paper 1 – 2,5 hours**

1. Exponents
2. Surds
3. Equations and inequalities, completing the square
4. Nature of roots
5. Number Patterns
6. Functions – straight line, parabola, hyperbola, exponential graphs
7. Finance
8. Probability – 3 circle Venn diagrams, tree diagrams, contingency tables, independent events

#### **Paper 2 – 2,5 hours**

1. Analytical geometry
2. Trigonometry – reduction formulae, equations, identities, trig on the Cartesian plane
3. Trig graphs
4. Circle Eudidean geometry
5. Statistics
6. Trigonometry – solving triangles using SOH CAH TOA, sin rule, cos rule and area rule
7. Measurement

### **Life Sciences**

#### **Paper 1 – 2,5 hours**

1. Supporting systems and movement in animals – human skeleton, joints, and muscles.
2. Transport in animals – the human circulatory system, heart structure, blood vessels, blood, and lymph as a tissue.
3. Excretion in humans and the urinary system.
4. The Human nervous system
  - The nervous systems
  - The brain, spinal cord, and reflex arc
  - Nerve tissue and transmission of nerve impulses
  - Receptors

# Grade 11 Mathematics Paper 1 Scope In November Exam

**Rosina Ehmann**



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subsequent  $\Delta h$  values  $2 \text{ CO}_2(\text{g}) + \text{H}_2\text{O}(\text{g}) \rightarrow \text{C}_2\text{H}_2(\text{g}) + \frac{5}{2} \text{O}_2(\text{g}) \quad \Delta h = 94.5 \text{ kJ}$   
 $\text{H}_2\text{O}(\text{g}) \rightarrow \text{H}_2(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \quad \Delta h = 71.2 \text{ kJ}$   
 $\text{C}_2\text{H}_6(\text{g}) + \frac{7}{2} \text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g}) + 3 \text{H}_2\text{O}(\text{g}) \quad \Delta h = 283 \text{ kJ}$   
 $\text{C}_2\text{H}_6(\text{g})$

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web chemistry 120 hess s law worksheet calculate  $\Delta h$  for the reaction  $\text{C}_2\text{H}_4(\text{g}) + 2 \text{Cl}_2(\text{g}) \rightarrow \text{C}_2\text{H}_2\text{Cl}_2(\text{g}) + 2 \text{HCl}(\text{g})$  from the following data  
 $\text{C}_2\text{H}_4(\text{g}) + 3 \text{O}_2(\text{g}) \rightarrow 2 \text{CO}_2(\text{g}) + 2 \text{H}_2\text{O}(\text{l}) \quad \Delta h = 1411 \text{ kJ/mole}$   
 $\text{C}_2\text{H}_2\text{Cl}_2(\text{g}) + 2 \text{O}_2(\text{g}) \rightarrow 2 \text{CO}_2(\text{g}) + \text{H}_2\text{O}(\text{l}) \quad \Delta h = 1560 \text{ kJ/mole}$   
 $\text{H}_2(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l}) \quad \Delta h = 285 \text{ kJ}$

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web thus for reaction 9 2 in which 1 mole of  $\text{H}_2\text{SO}_4$  and 1 mole of  $\text{NaOH}$  combine to form 2 moles of  $\text{H}_2\text{O}$   $\Delta H = 55.8 \text{ kJ}$  as shown note that the enthalpy change  $\Delta H_{\text{products}} - \Delta H_{\text{reactants}}$  is positive if heat is absorbed that is if  $\Delta H_{\text{products}}$

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web objective the purpose of the lab is to combine two reactants in the calorimeter and measure the heats of reaction in solution eventually allowing an individual to prove or disprove hess s law introduction thermodynamics is the study of energy exchanged between the system and the surroundings

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web 3 12 hess s law in this experiment you will measure the amount of heat released in these three related exothermic reactions  $\text{NaOH}(\text{s}) + \text{NaCl}(\text{aq}) \rightarrow \text{NaOH}(\text{aq}) + \text{NaCl}(\text{aq})$   $\text{H}_2\text{SO}_4(\text{l}) + \text{NaOH}(\text{aq}) \rightarrow \text{Na}_2\text{SO}_4(\text{aq}) + \text{H}_2\text{O}(\text{l})$   $\text{H}_2\text{SO}_4(\text{l}) + \text{NaCl}(\text{aq}) \rightarrow \text{Na}_2\text{SO}_4(\text{aq}) + \text{H}_2\text{O}(\text{l})$