

100) A student finds the ΔG of hydrolysis of ATP is -30.5 kJ/mol at pH 7.0. Is this value the same as the ΔG of phosphorylation? At which condition would phosphorylation be most favorable?

101) What does the term "coupled reaction" mean?

_____ ΔG of 1 reaction is used to drive ΔG of 2nd reaction _____ 200

102) What is the standard free energy change of the hydrolysis of ATP?

_____ -30.5 kJ/mol _____ 200

103) Calculate the ΔG of phosphorylation of ADP to ATP at pH 7.0.

$\frac{1.2 \times 10^{-3}}{2 \times 10^{-3}} = \frac{e^{-\Delta G/RT}}{e^{-\Delta G/RT}}$ $\Delta G = -RT \ln \left(\frac{[ADP][P_i]}{[ATP]} \right)$
 $\Delta G = - (8.314 \text{ J/mol}\cdot\text{K}) (298 \text{ K}) \ln \left(\frac{1.2 \times 10^{-3}}{2 \times 10^{-3}} \right)$
 $\Delta G = 1.2 \text{ kJ/mol}$ _____ 200

104) What is the standard free energy change of the hydrolysis of ATP at pH 7.0?

105) Explain why the hydrolysis of ATP is a coupled reaction.

_____ The hydrolysis of ATP is a coupled reaction
 _____ because it provides the energy to drive other reactions
 _____ that are not spontaneous _____ 200

106) Explain how a coupled reaction can be used to drive a non-spontaneous reaction.

_____ Explain what is meant by the term "coupled reaction" and suggest other ATP-coupled reactions.

_____ The hydrolysis of ATP is a coupled reaction
 _____ because it provides the energy to drive other reactions
 _____ that are not spontaneous _____ 200

(Total: 200)



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