



10. Which of the following are endothermic processes?

- A. ice melting
- B. water condensing
- C. freezing
- D. sublimation

11. Heat (enthalpy) is released in EXO, positive temperature change.

12. Temperature following is a statement of the first law of thermodynamics?

- A. is negative change in enthalpy corresponds to an exothermic process
- B. Energy lost by the system must be gained by the surroundings, neither of the two is created
- C. Temperature is a measure of average kinetic energy
- D. Enthalpy is a state function

13. The system enthalpy of a closed process,  $\Delta H$  is negative. This means that the product is \_\_\_\_\_

- A. Enthalpy is a state function, enthalpy is conserved
- B. Enthalpy and the temperature of the surroundings will decrease
- C. Enthalpy and the surroundings will increase
- D. Enthalpy and the temperature of the surroundings will increase
- E. None of these is a true statement

14. The reaction



is \_\_\_\_\_ and therefore heat is \_\_\_\_\_ from the system.

- A. Endothermic, released
- B. Exothermic, absorbed
- C. Endothermic, released
- D. Exothermic, absorbed
- E. Thermodynamic value released and absorbed

15. Heat evolved in this reaction when 10.0 g of  $\text{C}_2\text{H}_2\text{Cl}_2$  is completely consumed in the reaction  
 given below is  $\text{C}_2\text{H}_2\text{Cl}_2 + \text{H}_2\text{O} \rightarrow \text{C}_2\text{H}_4\text{Cl}_2 + \text{HCl}$   $\Delta H = -100 \text{ kJ}$



- A. 100 kJ
- B. 200 kJ
- C. 100 or 200 kJ
- D. 200 kJ
- E. 10 kJ



# Mcps Honors Chemistry Exam Guide

**Jicheng Xie**



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