

PROCEDURE:

Rinse and fill the burette with the given solution C-10 (KMnO_4). Pipette out 20 ml or 25 ml of C-11 (hydrated ammonium iron (II) sulphate solution) and transfer into a clean conical flask. To this, add 20 ml of C-12 (dilute sulphuric acid) specially provided for titration.

Titrate the solution in the conical flask with C-10 (KMnO_4) slowly, till one drop of this gives a light permanent pink colour to the solution C-11 in the flask. The pink colour should not disappear on shaking the contents in the conical flask.

Repeat the above procedure to get at least two concordant readings.

Tabulate your readings.

State:

- (a) The capacity of the pipette used.
- (b) The titre value you intend to use in your calculations.

Show the titre value to the Visiting Examiner.

The equations for the above reactions are as follows:



The ionic equation for the reaction is:



Relative atomic masses:



Calculate the following:

- (i) The molarity of the solution of potassium manganate(VII) C-10.
- (ii) The molarity of hydrated ammonium iron(II) sulphate solution C-11.
- (iii) The molecular mass of hydrated ammonium iron(II) sulphate deduced from the experimental data.
- (iv) The numerical value of x , i.e. the number of molecules of water of crystallization in $(\text{NH}_4)_2\text{SO}_4 \cdot \text{FeSO}_4 \cdot x\text{H}_2\text{O}$.

Note: Molarity must be calculated up to at least 4 decimal places.

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