

change of the surrounding water. $\Delta H_{\text{metal}} = -\Delta H_{\text{water}}$ If the water temperature increases, then the metal must have released heat. We can find the enthalpy change of the metal with the formula: $\Delta H_{\text{metal}} = -m_{\text{water}} \cdot C_{\text{water}} \cdot \Delta T_{\text{water}}$

Check-for-understanding question:

1. When 25 mL of water containing 0.85 mol HCl is mixed with 40 mL of water containing .05 mol KOH in a calorimeter (with initial temperature 23 °C), the temperature of the solution increases to 25.5 °C. How many kilojoules (kJ) of heat are released by the acid-base reaction?

$$\begin{aligned}\Delta H_{\text{water}} &= m_{\text{water}} \times C_{\text{water}} \times \Delta T_{\text{water}} \\ &= 65 \text{ g} \times 4.18 \text{ J/g}^\circ\text{C} \times 2.5^\circ\text{C} \\ &= 679.25 \text{ J} = .67925 \text{ kJ}\end{aligned}$$

Measuring And Expressing Enthalpy Changes Answers

Kaplan Test Prep

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In a digitally-driven world where displays reign great and immediate conversation drowns out the subtleties of language, the profound secrets and emotional nuances hidden within words frequently move unheard. However, located within the pages of **Measuring And Expressing Enthalpy Changes Answers** a captivating fictional prize pulsating with fresh thoughts, lies a fantastic quest waiting to be undertaken. Penned by an experienced wordsmith, this marvelous opus invites visitors on an introspective journey, lightly unraveling the veiled truths and profound impact resonating within the material of each and every word. Within the emotional depths of this poignant evaluation, we will embark upon a heartfelt exploration of the book's key subjects, dissect their interesting publishing fashion, and yield to the strong resonance it evokes deep within the recesses of readers' hearts.

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