

# Mark Scheme

## Electrolysis – Molten

1. Ions cannot move
2. Anode (Remember PANIC: Positive Anode Negative Is Cathode).
3. Elements
4. Electrolyte
5. OILRIG - Oxidation Is Loss, Reduction Is Gain (of electrons).
6. Oxidation
7.  $\text{Ca}^{2+}$  would go to the cathode,  $\text{I}^-$  would go to the anode.
8.  $\text{Ca}^{2+} + 2\text{e}^- \rightarrow \text{Ca}$
9.  $2\text{Cl}^- \rightarrow \text{Cl}_2 + 2\text{e}^-$
10.  $\text{CaCl}_2 \rightarrow \text{Ca}^{2+} + 2\text{Cl}^-$

## Electrolysis – Solutions

1. NaCl
2.  $\text{Na}^+$ ,  $\text{H}^+$ ,  $\text{OH}^-$ ,  $\text{Cl}^-$
3. Hydrogen, chlorine, sodium hydroxide.
4. Negative ions ( $\text{OH}^-$ ,  $\text{Cl}^-$ )
5. Positive Ions ( $\text{H}^+$ ,  $\text{Na}^+$ )
6. Chlorine,  $2\text{Cl}^- \rightarrow \text{Cl}_2(\text{g}) + 2\text{e}^-$
7. Hydrogen,  $2\text{H}^+ + 2\text{e}^- \rightarrow \text{H}_2(\text{g})$
8.  $\text{Na}^+$  and  $\text{OH}^-$
9. Sodium hydroxide
10. Aluminium. Oxygen is released at the anode where it reacts with the graphite to form carbon dioxide. Therefore the anode needs to be replaced often.

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