

13. Which type of bond is formed when a Lewis acid reacts with a Lewis base?

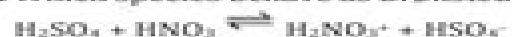
- A. Covalent
- B. Dipole-dipole
- C. Double
- D. Hydrogen

14. What is the order of increasing acidity of the following acids?

| Acid           | $K_a$                | Acid              | $pK_a$ |
|----------------|----------------------|-------------------|--------|
| chloroethanoic | $1.3 \times 10^{-3}$ | hydrogen fluoride | 3.3    |
| ethanoic       | $1.7 \times 10^{-5}$ | hydrogen cyanide  | 9.3    |

- A. chloroethanoic < ethanoic < hydrogen fluoride < hydrogen cyanide
- B. ethanoic < chloroethanoic < hydrogen fluoride < hydrogen cyanide
- C. chloroethanoic < ethanoic < hydrogen cyanide < hydrogen fluoride
- D. hydrogen cyanide < ethanoic < hydrogen fluoride < chloroethanoic

15. Which species behave as Brønsted-Lowry bases in the following reaction?



- A.  $\text{HNO}_3$  and  $\text{HSO}_4^-$
- B.  $\text{HNO}_3$  and  $\text{H}_2\text{NO}_3^+$
- C.  $\text{H}_2\text{SO}_4$  and  $\text{HSO}_4^-$
- D.  $\text{H}_2\text{NO}_3^+$  and  $\text{HSO}_4^-$

16. Which is a conjugate Brønsted-Lowry acid-base pair?



- A.  $\text{CH}_3\text{COO}^- / \text{H}_3\text{O}^+$
- B.  $\text{H}_2\text{O} / \text{CH}_3\text{COO}^-$
- C.  $\text{H}_2\text{O} / \text{H}_3\text{O}^+$
- D.  $\text{CH}_3\text{COOH} / \text{H}_2\text{O}$

17. Which of the following gases does **not** result in acid deposition?

- A.  $\text{CO}_2$
- B.  $\text{NO}_2$
- C.  $\text{NO}$
- D.  $\text{SO}_2$

18. What occurs when solid sodium hydrogen carbonate reacts with aqueous sulfuric acid?

- A. Bubbles of sulfur dioxide form.
- B. Bubbles of both hydrogen and carbon dioxide form.
- C. Bubbles of hydrogen form.
- D. Bubbles of carbon dioxide form.

19. A student carried out a titration to determine the concentration of an acid and found that his value had good precision but poor accuracy. Which process explains this outcome?

- A. Consistently overshooting the volume of solution from the burette into the flask.
- B. Collection of insufficient titration data.
- C. Reading the meniscus in the burette at a different angle each time.
- D. Forgetting to rinse the flask after one of the titrations.

# Ib Question Acids And Bases

**Sir William Grant MacPherson**



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