Modern Chemistry Chapter 14 Acids and Bases Notes from Reading

Section 1 Properties of Acids and Bases Pages 466-477

- List five properties of acids.
- 2. What is a binary acid?
- 3. How are the names of binary acids determined?
- 4. What is an oxyacid?
- 5. How are the names of oxyacids determined?
- 6. Read pages 470-471. In a table similar to the one below, write some properties of each acid.

H ₂ SO ₄	HNO ₃	H ₃ PO ₄	HCI	HC ₂ H ₃ O ₂
	7		- Harring	

- List five properties of bases.
- 8. What is the Arrhenius definition of an acid?
- 9. Compare and contrast the following: a proton, a hydrogen ion, a hydronium ion.
- 10. What is the Arrhenius definition of a base?
- 11. What makes an acid strong?
- 12. List seven strong acids.
- 13. What makes an acid weak?
- 14. Describe an organic acid and draw a structural formula for an example of an organic acid?
- 15. Read Chapter 18 pages 605-606. What is Ka?
- 16. How does Ka relate to the strength of an acid?
- 17. Compare and contrast a strong base and a weak base.
- 18. By the chemical formula how can a strong base be discerned from a weak base?

Section 2 Acid-Base Theories 478-482

- 19. What is the Bronsted-Lowry definition of an acid?
- 20. What is the Bronsted-Lowry definition of a base?
- Give an example of an equilibrium system that contains a Bronsted-Lowry acid and a Bronsted-Lowry base.
- 22. What is a monoprotic acid? a polyprotic acid? a diprotic acid? a triprotic acid?
- 23. Are organic acid monoprotic or polyprotic? Give an reason to support your answer.
- 24. What is the Lewis definition of an acid?
- 25. What is the Lewis definition of a base?

Section 3 Acid-Base Theories 483-899

- 26. What is a conjugate base?
- 27. What is a conjugate acid?
- Five an equilibrium system that illustrates a conjugate acid-base pair.
- 29. What is an amphoteric compound? Give an example.
- 30. What is a neutralization reaction?
- 31. Give a generic form of a neutralization reaction: + → +
- 32. What is the definition of a salt?

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