

- Q.65** The nitrogenous base different in RNA as compared to DNA is
 A) Cytosine
 B) Adenine
 C) Thymine
 D) Guanine
- Q.66** Carbohydrates are stored in human body as the polysaccharide
 A) Glycogen
 B) Starch
 C) Glucose
 D) Galactase
- Q.67** Denaturation of protein is used by
 A) Addition of sulphuric acid
 B) Addition of dichromate ion
 C) Changing the pH
 D) All of these
- Q.68** In the reaction, $4\text{NH}_3 + 5\text{O}_2 \rightarrow 4\text{NO} + 6\text{H}_2\text{O}$, when one mole of ammonia and one mole of oxygen are made to react to completion, then
 A) 1.0 mol of H_2O is produced
 B) 1.5 mol of NO is formed
 C) All the oxygen is consumed
 D) All the ammonia is consumed
- Q.69** In which set all isotopes have maximum natural abundance
 A) $^{40}_{20}\text{Ca}$, $^{12}_6\text{C}$, $^{35}_{17}\text{Cl}$
 B) $^{40}_{20}\text{Ca}$, $^{56}_{26}\text{Fe}$, $^{24}_{12}\text{Mg}$
 C) $^{40}_{20}\text{Ca}$, $^{23}_{11}\text{Na}$, $^{37}_{17}\text{Cl}$
 D) $^{20}_{10}\text{Ne}$, $^{32}_{16}\text{S}$, $^{14}_7\text{N}$
- Q.70** Two balloons are filled with equal moles of hydrogen and helium gas at STP. Which balloon will contract first if same size holes are made in both of them
 A) Both will contract at same time
 B) Balloon filled with hydrogen gas
 C) Balloon filled with Helium gas
 D) Only hydrogen filled balloon will contract
- Q.71** In all living organisms, nucleic acids are responsible to their ability
 A) To reproduce and store genetic information
 B) To transmit genetic information
 C) To undergo mutation
 D) All of these
- Q.72** Ozone acts as pollutant when present in
 A) Troposphere
 B) Stratosphere
 C) Mesosphere
 D) Thermosphere
- Q.73** A primary pollutant is that
 A) Which causes cancer
 B) Is waste product of an industry
 C) Whose concentration cannot be controlled
 D) Which corrodes the metals only
- Q.74** The pH of a neutral sample of water at 100°C when $K_w = 1.0 \times 10^{-12}$ is
 A) 7
 B) 6
 C) 0
 D) 12
- Q.75** The two steps involved in the gas phase reaction $\text{X} + 2\text{Y} \rightarrow \text{XY}_2$ are shown below
 $\text{X} + \text{Y} \xrightarrow{\text{slow}} \text{XY}$, $\text{XY} + \text{Y} \xrightarrow{\text{fast}} \text{XY}_2$
 What is the rate equation for the overall reaction
 A) Rate = $[\text{x}]^1[\text{y}]^1$
 B) Rate = $[\text{x}]^1[\text{y}]^2$
 C) Rate = $[\text{x}]^0[\text{y}]^2$
 D) Rate = $[\text{xy}]^1[\text{y}]^1$
- Q.76** Which of the following statements about the reaction given below is/are correct
 $\text{IO}_3^-(\text{aq}) + 2\text{I}^-(\text{aq}) + 6\text{H}^+(\text{aq}) + 6\text{Cl}^-(\text{aq}) \rightarrow 3\text{ICl}_2(\text{aq}) + 3\text{H}_2\text{O}(\text{l})$
 A) The oxidation number of iodine in the iodate ion, $\text{IO}_3^-(\text{aq})$, changes from +5 to +1
 B) The oxidation number of iodine in the iodide ion I^- , changes from -1 to +2
 C) The oxidation number of chlorine in the iodide ion I^- , changes from -1 to -2
 D) Iodate ions are oxidized
- Q.77** The one having highest ionization energy is
 A) Mg
 B) Al
 C) S
 D) P
- Q.78** The electronic configuration of elements across each period in the periodic table starts with a s-subshell of the new shell and ends with
 A) Small subshell
 B) The next higher shell
 C) Different subshell of the same subshell
 D) p-subshell of the same shell
- Q.79** Why binary compounds of oxygen and fluorine are called fluorides rather than oxides
 A) They always contain F-ions
 B) O atom is larger than F atom
 C) F is more electronegative than O
 D) O is better oxidizing agent
- Q.80** All the elements of group IVA do not have which of the property in common
 A) Show four valency
 B) Form tetrahalide
 C) Show inert pair effect
 D) Form dioxide

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Table of Contents Mcat Full Length 1 Answers

1. Understanding the eBook Mcat Full Length 1 Answers
 - The Rise of Digital Reading Mcat Full Length 1 Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Mcat Full Length 1 Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Mcat Full Length 1 Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Mcat Full Length 1 Answers
 - Personalized Recommendations
 - Mcat Full Length 1 Answers User Reviews and Ratings
 - Mcat Full Length 1 Answers and Bestseller Lists

5. Accessing Mcat Full Length 1 Answers Free and Paid eBooks
 - Mcat Full Length 1 Answers Public Domain eBooks
 - Mcat Full Length 1 Answers eBook Subscription Services
 - Mcat Full Length 1 Answers Budget-Friendly Options
6. Navigating Mcat Full Length 1 Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Mcat Full Length 1 Answers Compatibility with Devices
 - Mcat Full Length 1 Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Mcat Full Length 1 Answers
 - Highlighting and Note-Taking Mcat Full Length 1 Answers
 - Interactive Elements Mcat Full Length 1 Answers
8. Staying Engaged with Mcat Full Length 1 Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Mcat Full Length 1 Answers
9. Balancing eBooks and Physical Books Mcat Full Length 1 Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Mcat Full Length 1 Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Mcat Full Length 1 Answers
 - Setting Reading Goals Mcat Full Length 1 Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Mcat Full Length 1 Answers
 - Fact-Checking eBook Content of Mcat Full Length 1 Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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