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Shelby Johnson



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web nov 23 2017 maharashtra x physics suppose the masses of calorimeter th water in it and the hot object made up of copper which is put in the calorimeter are the same the initial temperature of the calorimeter and water is 30 c and that of hot object is 60 c the specific heat of copper and water are 0 09 cal gm c and 1 cal gm c respectively

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web oct 28 2019 solution use this equation $q = \text{specific heat} \times m \times \Delta t$ where q is heat flow m is mass in grams and Δt is the temperature change plugging in the values given in the problem you get $q_{\text{water}} = 4.18 \text{ J/g} \cdot \text{C} \times 110 \text{ g} \times 26.6 \text{ C} = 12100 \text{ J}$ $q_{\text{water}} = 550 \text{ J}$ $\Delta h_{\text{q water}} = 550 \text{ J}$

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web calorimetry problems calorimetry problems qsur $m \times c \times t$ q heat m mass q_{rxn} q_{sur} t_{f} t_{i} name per date c specific heat for water 4.184 J/g°C 1 what is the specific heat of aluminum if the temperature of a 28.4 g sample of aluminum is increased by 8.1 °C when 207 J of heat is added 2

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web jul 12 2023 this concept lies at the heart of all calorimetry problems and calculations because the heat released or absorbed at constant pressure is equal to ΔH the relationship between heat and ΔH_{rxn} is $\Delta H_{\text{rxn}} = q_{\text{rxn}} / n$ $q_{\text{calorimeter}} = m c \Delta t$ 7 3 14 7 3 14 $\Delta H_{\text{rxn}} = q_{\text{rxn}} / n$ $q_{\text{calorimeter}} = m c \Delta t$

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web use the information in this chart to answer the following questions 1 if i have 125 grams of water at a temperature of 25.0 °C and heat it to 145.0 °C what will the change in enthalpy be 2 sketch the heating curve temperature vs time for the process in 1 3 i have 18.7 grams of an unknown liquid chemical at a temperature of 35.0 °C

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web sample problem calorimetry and enthalpy changes in an experiment 25.0 mL of 1.00 M HCl at 25.0 °C is added to 25.0 mL of 1.00 M NaOH at 25.0 °C in a foam cup calorimeter a reaction occurs and the temperature rises to 32.0 °C calculate the enthalpy change in kJ for this reaction assume the densities of the solutions are 1.00 g/mL and that

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web before we practice calorimetry problems involving chemical reactions consider a simple example that illustrates the core idea behind calorimetry suppose we initially have a high temperature substance such as a hot piece of metal m and a low temperature substance such as cool water w answer the initial temperature of the copper

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web choose 1 answer the final temperature is closer to t 1 than to t 2 a the final temperature is closer to t 1 than to t 2 the final temperature is exactly halfway between t 1 and t 2 b the final temperature is exactly halfway between t 1 and t 2 the final temperature is closer to t 2 than to t 1 c

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web in this set of practice questions we will go over the main types of questions on calorimetry including the heat capacity the heat of reaction finding the final temperature of a mixture constant pressure calorimetry and constant volume calorimetry a few important concepts and formulas you will need to solve these questions

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web calorimetry worksheet 1 if 0 315 moles of hexane c₆h₁₄ is combusted in a bomb calorimeter containing 5 65 liters of water calculate the molar heat of combustion of hexane if the water temperature rises 55 4 c the specific heat capacity of water is 4 184 j g c h ms t h 5 650 grams h₂o 4 184 j g c 55 4 c h 1310 kj

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web 6 a 218 g sample of steam at 121oc is cooled to ice at 14oc find the change in heat content of the system 7 in going from ice at 34oc to steam at 138oc a sample of water absorbs 1 41 x 10⁵ j find the mass of the sample answers 1 2 68 x 10⁴ j 2 6

23 x 104 j 3 1 71 x 104 j 4 1 11 x 105 j 5 2 11 x 105 j 6

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