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For the chemical reaction

Chemical	State	Stoichiometric Coefficient	Standard Enthalpy of Formation, ΔH_f° (kJ/mol)	Standard Enthalpy of Combustion, ΔH_c° (kJ/mol)
H_2	(g)	2	0	0
O_2	(g)	1	0	0
H_2O	(l)	2	-285.8	-285.8

The grams of water that can be produced from 2.192g of hydrogen can be calculated as follows.

$(2.192 \text{ g H}_2) \left(\frac{1 \text{ mol H}_2}{2.016 \text{ g H}_2} \right) \left(\frac{2 \text{ mol H}_2\text{O}}{2 \text{ mol H}_2} \right) \left(\frac{18.015 \text{ g H}_2\text{O}}{1 \text{ mol H}_2\text{O}} \right)$

$(2.192 \text{ g}) \left(\frac{1 \text{ mol H}_2\text{O}}{2.016} \right)$

Guide To Microsoft Equation Editor 3.0

Michal Rosen-Zvi



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