

## Question 1

- 1(b)(i) Reaction rate decreased as time increased.
- 1(b)(ii) The smaller zinc particles gave a larger surface area which led to a faster reaction rate since there were more successful collisions per unit time.
- 1(b)(iii) Temperature, pressure (for gases) and concentration (of reagents in solution).
- 1(c) One safety precaution could have been the wearing of safety glasses/goggles to protect the eyes from possible acid spray.
- 1(d) Temperature could have been controlled during the experiment by the use of a water bath.
- 1(e)(i)  $\text{Zn}_{(s)} + \text{H}_2\text{SO}_{4(aq)} \rightarrow \text{ZnSO}_{4(aq)} + \text{H}_{2(g)}$
- 1(e)(ii) From the balanced equation for the reaction, 1 mole of  $\text{Zn}_{(s)}$  produces 1 mole of  $\text{H}_{2(g)}$ .  
 Moles of zinc used = mass of Zn used/molar mass of Zn =  $(0.55/65.4)$  mol, i.e. 0.0084 mol. Zn.  
 Therefore 0.0084 moles of  $\text{H}_{2(g)}$  were produced.  
 Volume of  $\text{H}_{2(g)}$  produced = moles of  $\text{H}_{2(g)}$  produced  $\times$  molar volume of gas at RTP.  
 Volume of  $\text{H}_{2(g)}$  produced =  $(0.0084 \times 24\,000) \text{ cm}^3 = 201.8 \text{ cm}^3$ .

## Question 2

- 2(a) The number of occupied electron shells in an tells us what period it is in. All these elements have three occupied electron shells and are therefore in period three.

- 2(b) Element: Sodium (Na).



- 2(c) Allotrope – one of multiple (2 or more) different forms/structures of the same element, which exist in the same physical state.

- 2(d)     
 × – (sulfur) electron.  
 S – nucleus of sulfur atom.

- 2(e)(i) When sulfur reacts with chlorine there is covalent bonding.

- 2(e)(ii) Both sulfur and chlorine are non-metals. It is energetically preferable for these atoms to share electrons with each other in order to achieve stable outer electronic configurations of 8 electrons, rather than to lose or gain electrons.

- 2(f)  $\text{SCl}_2$ .

- 2(g)     
 S – sulfur nucleus and inner electrons, × – sulfur electron,  
 Cl – chlorine nucleus and inner electrons, • – chlorine electron.

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**Ying-Ying Zheng**



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