

## Mark Scheme

Q1.

| Question Number | Acceptable Answer                                                                                                                                                                                          | Additional Guidance                                                                                                                                                                                                                                                                                                        | Mark |
|-----------------|------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------|----------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------|------|
| (a) (i)         | $\text{CaCO}_3(\text{s}) + 2\text{HCl}(\text{aq}) \rightarrow \text{CaCl}_2(\text{aq}) + \text{H}_2\text{O}(\text{l}) + \text{CO}_2(\text{g})$<br><br>Balanced equation<br>(1)<br><br>State symbols<br>(1) | Accept<br>$\text{CaCO}_3(\text{s}) + 2\text{H}^+(\text{aq}) \rightarrow \text{Ca}^{2+}(\text{aq}) + \text{H}_2\text{O}(\text{l}) + \text{CO}_2(\text{g})$<br><br>2nd mark dependent on first or near miss.<br><br>Reject $\text{H}_2\text{CO}_3(\text{aq})$ in equation, but allow state symbol mark if otherwise correct. | (2)  |
| Question Number | Acceptable Answer                                                                                                                                                                                          | Additional Guidance                                                                                                                                                                                                                                                                                                        | Mark |
| (a) (ii)        | Finds molar mass of calcium carbonate                                                                                                                                                                      | Example of calculation<br>$\text{Mr of calcium carbonate}$<br>$= 40.1 + 12 + (16 \times 3) = 100.1 \text{ (g mol}^{-1}\text{)}$<br>Allow<br>$= 40 + 12 + (16 \times 3) = 100 \text{ (g mol}^{-1}\text{)}$<br>Accept answer with no working                                                                                 | (1)  |

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**AW Rasmussen**



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