

11. In the reaction:  $4\text{NH}_3 + 5\text{O}_2 \rightarrow 4\text{NO} + 6\text{H}_2\text{O}$ , if 20 grams of  $\text{NH}_3$  and 50 grams of  $\text{O}_2$  are mixed, which is the limiting reactant, and how many grams of  $\text{NO}$  can be produced? (Molar masses:  $\text{NH}_3 = 17 \text{ g/mol}$ ,  $\text{O}_2 = 32 \text{ g/mol}$ ,  $\text{NO} = 30 \text{ g/mol}$ )
12. For the reaction:  $2\text{Al} + 3\text{H}_2\text{SO}_4 \rightarrow \text{Al}_2(\text{SO}_4)_3 + 3\text{H}_2$ , if 10 grams of  $\text{Al}$  and 50 grams of  $\text{H}_2\text{SO}_4$  are reacted, which reactant is limiting, and how many grams of  $\text{H}_2$  can be produced? (Molar masses:  $\text{Al} = 27 \text{ g/mol}$ ,  $\text{H}_2\text{SO}_4 = 98 \text{ g/mol}$ ,  $\text{H}_2 = 2 \text{ g/mol}$ )
13. In the reaction:  $3\text{CuSO}_4 + 2\text{Na}_3\text{PO}_4 \rightarrow \text{Cu}_3(\text{PO}_4)_2 + 3\text{Na}_2\text{SO}_4$ , if 30 grams of  $\text{CuSO}_4$  and 20 grams of  $\text{Na}_3\text{PO}_4$  are mixed, which is the limiting reactant, and how many grams of  $\text{Cu}_3(\text{PO}_4)_2$  can be produced? (Molar masses:  $\text{CuSO}_4 = 160 \text{ g/mol}$ ,  $\text{Na}_3\text{PO}_4 = 164 \text{ g/mol}$ ,  $\text{Cu}_3(\text{PO}_4)_2 = 382 \text{ g/mol}$ )
14. For the reaction:  $2\text{KMnO}_4 + 5\text{H}_2\text{C}_2\text{O}_4 + 3\text{H}_2\text{SO}_4 \rightarrow 2\text{MnSO}_4 + 10\text{CO}_2 + 8\text{H}_2\text{O} + \text{K}_2\text{SO}_4$ , if 20 grams of  $\text{KMnO}_4$ , 30 grams of  $\text{H}_2\text{C}_2\text{O}_4$ , and 25 grams of  $\text{H}_2\text{SO}_4$  are reacted, which reactant is limiting, and how many grams of  $\text{CO}_2$  can be produced? (Molar masses:  $\text{KMnO}_4 = 158 \text{ g/mol}$ ,  $\text{H}_2\text{C}_2\text{O}_4 = 90 \text{ g/mol}$ ,  $\text{H}_2\text{SO}_4 = 98 \text{ g/mol}$ ,  $\text{CO}_2 = 44 \text{ g/mol}$ )

# Holt Chemistry Limiting Reactants And Percentage Yield

**SB Merriam**



## **Holt Chemistry Limiting Reactants And Percentage Yield:**

## Reviewing **Holt Chemistry Limiting Reactants And Percentage Yield**: Unlocking the Spellbinding Force of Linguistics

In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is truly astonishing. Within the pages of "**Holt Chemistry Limiting Reactants And Percentage Yield**," an enthralling opus penned by a highly acclaimed wordsmith, readers embark on an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve to the book is central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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