

An Instructor's Solution Manual to Accompany  
Engineering Mechanics: Dynamics, 3<sup>rd</sup> Edition

Andrew Pytel and Jaan Kiusalaas



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# Mechanics Dynamics Solution Manual

**Andrew Pytel, Jaan Kiusalaas**



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web the equation below is the final net ionic equation after elimination of the spectator ions which yields water  $\text{h}^+\text{aq} + \text{oh}^-\text{aq} \rightarrow \text{h}_2\text{o}$  the spectator ions in  $\text{koh}^+\text{aq}$  and  $\text{hno}_3^-\text{aq}$  are  $\text{k}^+\text{aq}$  and  $\text{no}_3^-\text{aq}$  how to identify and list the

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web the  $\text{ag}^+\text{aq}$  and  $\text{cl}^-\text{aq}$  ions become  $\text{agcl}$  s but the  $\text{na}^+\text{aq}$  ions and the  $\text{no}_3^-\text{aq}$  ions stay as  $\text{na}^+\text{aq}$  ions and  $\text{no}_3^-\text{aq}$  ions these two ions are examples of spectator ions ions that do nothing in the overall course of a chemical reaction

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web jun 2 2023 the spectator ions are cut out from a reaction hence the net ionic equation for the above reaction would be  $2\text{Cl}^- + \text{Cu}^{2+} + 2\text{Cl}^- + \text{Cu}^{2+} \rightarrow 2\text{Cl}^- + \text{Cu}^{2+} + 2\text{Cl}^- + \text{Cu}^{2+}$  these spectator ions are not actively involved in the reaction the existence of spectator ions affects the debye length

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