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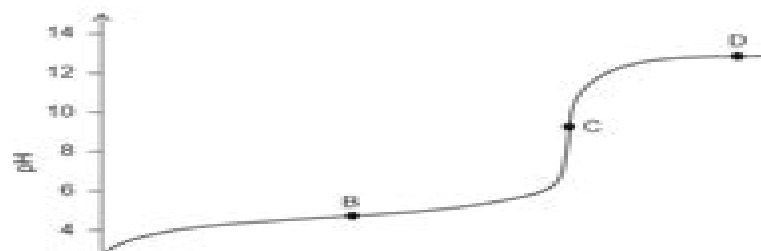
HL Paper 2

Benzoic acid, $\text{C}_6\text{H}_5\text{COOH}$, is another derivative of benzene.

- Identify the wavenumber of one peak in the IR spectrum of benzoic acid, using section 26 of the data booklet. [1]
 - Identify the spectroscopic technique that is used to measure the bond lengths in solid benzoic acid. [1]
 - Outline one piece of physical evidence for the structure of the benzene ring. [1]
 - Draw the structure of the conjugate base of benzoic acid showing **all** the atoms and **all** the bonds. [1]
 - Outline why both C to O bonds in the conjugate base are the same length and suggest a value for them. Use section 10 of the data booklet. [2]
- f(i). The pH of an aqueous solution of benzoic acid at 298 K is 2.95. Determine the concentration of hydroxide ions in the solution, using section 2 of the data booklet. [2]
- f(ii) Formulate the equation for the complete combustion of benzoic acid in oxygen using only integer coefficients. [2]
- The combustion reaction in (f)(ii) can also be classed as redox. Identify the atom that is oxidized and the atom that is reduced. [1]
 - Suggest how benzoic acid, $M_r = 122.13$, forms an apparent dimer, $M_r = 244.26$, when dissolved in a non-polar solvent such as hexane. [1]
 - State the reagent used to convert benzoic acid to phenylmethanol (benzyl alcohol), $\text{C}_6\text{H}_5\text{CH}_2\text{OH}$. [1]

A student performs a titration to determine the concentration of ethanoic acid, CH_3COOH , in vinegar using potassium hydroxide.

The pH curve for the reaction is given.



May 2014 Paper 1b Chem HL

Thomas Griffiths



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